

First Principles Study of Nickel Complex with 1,3-dithiole-2-thione-4,5-dithiolate Ligands as Model Photosensitizers

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Abstract

Dye-sensitized solar cells (DSSCs) have become in one important and promising technology in the photovoltaic field. The ability for a sensitizer to harvest light photons and inject the excited electrons into a photoanode, typically a metal oxide, determines the performance and operation range of the solar cell. Metal complexes with 1,3-dithiole-2-thione-4,5-dithiolate (dmit) ligands, which are an important class of functional materials, have received extensive attention due to their intriguing chemical and physical properties. The electronic and molecular properties of isolated and adsorbed nickel complexes with dmit ligands have been investigated using first principles calculations based on the density functional theory (DFT). Adsorption energies of metal complexes supported on the anatase TiO₂(101) surface were calculated for three different configurations, linked by sulphur atom of S_{thione}, S_{thiole-S_{thiolate}}, and planar. The most stable adsorption configurations found in this study are the S_{thiole-S_{thiolate}} and the planar forms for the nickel complex. TD-DFT molecular calculations reveal that the lowest energy transition in ultraviolet visible near-infrared (UV-Vis-NIR) mainly corresponds to the *HOMO-LUMO* $\pi-\pi^*$ excitation for the nickel complex. The effect of the TiO₂(101) surface on the absorption spectra of the nickel complex is practically limited to a red shift of about 0.1-0.3 eV. The analysis of the density of states for the dmit/TiO₂(101) system shows that the *LUMO* of the metal complex lies at the edge of the TiO₂ conduction band indicating, therefore, that electron injection from the complex excited state into the semiconductor surface is unlikely.

Keywords: Metal complexes, dmit ligand, DFT, TD-DFT, Adsorption energy, Sensitizer

Introduction

The sensitization of wide band-gap semiconductors lays at the heart of dye sensitized solar cells (DSSCs), a rising technology for solar energy harvesting that offers some advantages over classical Si-based devices [1, 2]. A key strength of DSSCs is the separation of electron generation and transport processes in two distinct materials. This allows the disconnected optimization of the dye for photon absorption and of a wide band-gap semiconductor for electron-hole separation and collection [3].

Typically metal oxides like zinc oxide (ZnO) [4, 5], stannic oxide (SnO₂) [6, 7, 8] and titanium dioxide (TiO₂) [2] have been used as the semiconductor material. However, different experimental results have shown that TiO₂ is preferable over either ZnO or SnO₂. Titanium dioxide is non-toxic, highly abundant, and provides a mesoporous structure for both organic and inorganic dye adsorption.

The role of the dye is to absorb the incoming photons and to transfer the excited electron to the conduction band of the semiconductor. Thus, an efficient dye should (a) be strongly adsorbed at the semiconductor surface; (b) show intense absorption in the visible and near infrared regions of the electromagnetic spectrum; (c) be stable enough as to be capable of multiple oxidation-reduction cycles and (d) be stable enough in its oxidized form as to be reduced by the electrolyte and its lowest excited state should be higher than the semiconductor conduction band edge.

Basically, photosensitizer dyes are either pure organic compounds or metal-based organometallic complexes. Metal-free organic sensitizers are cheaper, easy to modify structurally to tune the dye properties, in some cases they are environmentally benign and non-toxic and have high molar absorption coefficients. However, they also show important stability and efficiency problems. Many metal complexes-based dyes have been proposed. Of these, ruthenium (II) polypyridyl complexes have been shown to be the best so far [9]. However, the low abundance of the metal, its high cost and toxicity impose severe limitations on its practical and widespread use. Thus, transition metal complexes based on iron (Fe), nickel (Ni), cobalt (Co), palladium (Pd), platinum (Pt), and zinc (Zn), among others, have been proposed as alternatives in the design of photovoltaic sensitizers [2,3,10,11,12,13,14].

In this regard, square-planar complexes with sulfur-containing ligands that absorb in the near infrared region (NIR) of the spectrum have attracted special interest and have been examined both experimentally and theoretically [12,15,16,17,18,19]. Islam et al. [15] were the first to explore the application of a series of square-planar

diimine-dithiolate complexes as sensitizers. They synthesized and characterized a series of platinum-based polypyridyl complexes with dithiolate ligands that were also anchored to nanocrystalline TiO_2 in photoelectrochemical cells. The intense charge transfer band in these complexes was shown to be tunable by changing the dithiolate ligands. Geary et al. [16] prepared and examined a family of Pt(II)(diimine)(dithiolate) complexes, analyzing the influence of 3,3'-, 4,4'-, and 5,5'- bipyridyl substituents on their electronic properties. All synthesized complexes were attached to a TiO_2 substrate and tested as solar cells sensitizers with the 3,3'-disubstituted bipyridyl complex showing the highest photovoltaic performance.

In a later study [17] the superior performance of the 3,3'- bipyridyl complex was rationalized by using density functional theory calculations based on a hybrid functional that suggested that the longer-lived charge-separated state for this complex on TiO_2 was related to the non-planar geometry of the complex, reducing the electronic coupling between ligands. Lazarides et al. [18] have attempted to increase the light absorption properties of Pt(II)(diimine)(dithiolate) chromophores by combining them with boron-dipyrromethene, a strongly absorbing dye, in a dual chromophore system. By using time-dependent DFT calculations, the authors show that the many paths for electron transfer that exist in these systems result in unexpected routes for excited-state relaxation and loss of the desired properties of the excited charge transfer state. Despite the intense work developed on examining the potential of Pt(II)(diimine)(dithiolate) complexes as sensitizers for DSSC cells, only the paper by Linfoot et al. [12] has studied some Ni(II)(diimine)(dithiolate) dyes in relation with their use as dyes in a DSSC cell. The authors characterized the complexes using electrochemical, spectroscopic and computational techniques and assigned intense visible absorptions to ligand-to-ligand charge transfer transitions that would suggest appropriate charge separation for using on a photoelectrochemical device. However, low photocurrents were found when the complex was adsorbed on a TiO_2 film, a problem that was linked to a short-lived excited state of the Ni(II) complex.

Because of their unique properties related to applications in fields as diverse as conducting and superconducting materials, non-linear optics, catalysis, and dyes, metal dithiolene complexes, $\text{R}_2\text{M}(\text{dmit})_2$, $\text{R}=\text{PyMe}$, NEt_4 , NMe_4 , NPr_4 , NBu_4 , and $\text{dmit}=1,3$ -dithiole-2-thione-4,5-dithiolate) have been extensively studied for more than forty years [20]. These applications result from an interplay of different properties, including highly delocalized frontier orbitals that allow direct electron transfer through the ligand

π orbitals. For this reason, these complexes are considered promising candidates for photochemical devices [21].

Here, we present a study of the structural and spectroscopic properties of the model dithiolene complex $[(\text{CH}_3)_2][\text{Ni}(\text{dmit})_2]$ by combining DFT and TD-DFT calculations. First we analyze the properties of the isolated complex using the B3LYP functional and an atom-centered basis set. Second, the geometric and electronic properties of the complex adsorbed on a model TiO_2 (101) anatase surface have been examined by using plane-wave calculations that include both the use of a Hubbard correction to properly localize the metal d-electrons and an approximate functional to improve the description of the dispersion forces on the DFT calculations. Finally, we theoretically examine the performance of the model $[(\text{CH}_3)_2][\text{Ni}(\text{dmit})_2]$ complex as a sensitizer.

Computational details

For the isolated $[(\text{CH}_3)_2][\text{Ni}(\text{dmit})_2]$ complex, DFT calculations have been performed using the Gaussian 09 quantum chemical package [22]. Equilibrium geometry and electronic properties were determined by employing the hybrid Becke three-parameter functional with the Lee, Yang, and Parr (B3LYP) exchange correlation functional [23,24], with CEP-121G [25,26,27] effective core potentials and basis sets for Ni and S atoms and 6-31++G basis set for C and H atoms. Given that some charge-transfer character was found in the first excited state, the CAM-B3LYP hybrid functional, that includes long-range corrections was also employed [28]. TD-DFT single-point energy calculations were performed on optimized geometries.

To determine the geometric and electronic properties of the TiO_2 (101) surface and TiO_2 (101) surface with the adsorbed nickel complex, we performed periodic DFT calculations using the Vienna *ab initio* Simulation Package (VASP) [29,30,31]. The projector augmented wave (PAW) method was used, and the cutoff energy was set to 400 eV for slab and adsorption calculations. The generalized gradient approximation (GGA) functional was used (Perdew-Burke-Ernzerhof, PBE) [32]. In order to better render the anatase band gap, usually underestimated in plain GGA DFT calculations, a Hubbard type on-site Coulomb correction term was used as implemented by Dudarev et al. [33]. The GGA+ U procedure was applied on the transition metal d electrons, being the U_{eff} values employed in this work 4.5 eV and 5.5 eV for $3d$ levels of Ti and Ni,

respectively [34,35]. Optical spectra were obtained from the frequency dependent dielectrical functions as proposed by Gajdoš *et al.* [36].

Because GGA functionals neglect attractive long-range contributions, computed adsorption energies are generally underestimated [37]. To include the van der Waals corrections into the density functional approach (vdW-DF) and obtain a more accurate description, the method proposed by Tkatchenko and Scheffler was employed in this work [38].

The slab model of anatase surface was obtained by appropriately cutting the most stable TiO₂ (101) surface, and is represented by 96 [TiO₂] units arranged according to anatase crystalline structure. The model 5x3 supercell consisted of two O-Ti-O trilayers, 144 atoms each, where the bottom layer was fixed. The orthorhombic supercell has, thus, dimensions: $a = 31.254 \text{ \AA}$, $b = 15.288 \text{ \AA}$ and $c = 35.916 \text{ \AA}$, including a vacuum space of 20 \AA in the c direction. All calculations were performed at the Γ point.

Adsorption of [CH₃]₂[Ni(dmit)₂] complex on the TiO₂ (101) surface was done in three different adsorption configurations: linked by S_{thione}, linked by S_{thiole}-S_{thiolate} (bridge) and plane (Fig. 1).

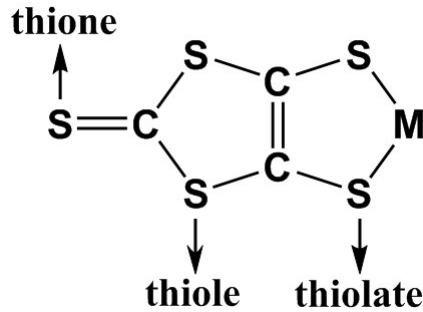


Fig. 1: Schematic structure of adsorption form S_{thione}, S_{thiole}-S_{thiolate}.

Adsorption energies (E_{ADS}) for the optimized metal complexes on the TiO₂(101) surface were calculated using

$$E_{ADS} = E_{(TiO_2)+(Ni_{dmit})} - (E_{(TiO_2)} + E_{Ni_{dmit}})$$

where $E_{(TiO_2)+(Ni_{dmit})}$ is the energy of (Ni-dmit) complex adsorbed on the TiO₂ (101) surface, $E_{Ni_{dmit}}$ and $E_{(TiO_2)}$ are the energies of the isolated Ni-dmit complex and clean TiO₂ (101) surface respectively. With this definition, negative adsorption energies represent bound states stable with respect to desorption.

Results and discussion

Structure and Electronic Properties of [CH₃]₂[Ni(dmit)₂] complex

Fig. 2 shows the optimized structure of the square planar complex [CH₃]₂[Ni(dmit)₂]. Table 1 presents the geometric parameters obtained at the B3LYP and PBE+U levels, in comparison to the experimental structure. We report only the relevant bond lengths and bond angles.

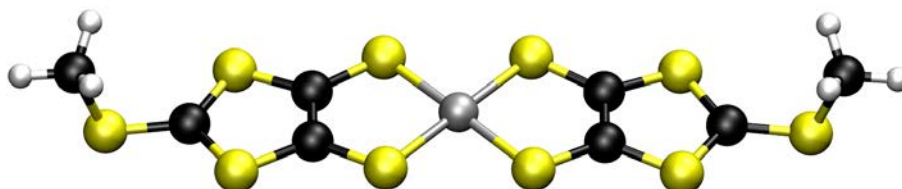


Fig. 2: Optimized structure of [CH₃]₂[Ni(dmit)₂] complex. Atoms colors code: Ni, gray; C, black; S, yellow; H, white.

Table 1: Main geometrical parameters calculated for [CH₃]₂[Ni(dmit)₂] complexes

	B3LYP	CAM-B3LYP	PBE+U	Exp [39]
<i>Bond distances</i>				
Ni-S	2.244	2.230	2.186	2.16-2.17
S=C	1.747	1.749	1.721	1.66
C=C	1.401	1.400	1.416	1.39
<i>Bond angles</i>				
S-Ni-S	92.6	92.8	93.1	92.2
S-Ni-S	87.4	87.2	93.2	86.6
Ni-S-C	101.5	101.4	102.4	102.8

*Bond lengths in Å and bond angles in degrees

From the data shown in Table 1, a general agreement between calculated and experimental values is observed. Optimized bond distances are systematically overestimated, and the values obtained from PBE+U calculations in general are in better

agreement than those estimated with either the B3LYP or the CAM-B3LYP functionals, except for the C=C double bond. Ni–S bond lengths are very similar to each other and in agreement with the experimental results. In contrast, the S=C bond is significantly overestimated. The disagreement found can be related to the fact that experimental data derive from solid crystal structure diffraction experiments in which packing forces may alter the geometry of individual molecules.

The calculated harmonic vibrational frequencies and band assignments for the nickel complex are presented in Table 2.

Table 2: Comparison between the experimental and calculated frequencies and assignments of vibrational modes of $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$ (cm^{-1})

	B3LYP	CAM-B3LYP	PBE+U	Exp [40,41,42,43,44]
$\nu_s(\text{C-H})$	3073	3098	3001	3000
$\nu_s(\text{C=C})$	1331	1374	1297	1454
$\nu(\text{C-S}_{\text{thiole}})$	943	984	933	940
$\nu(\text{S}-(\text{C}=\text{S}_{\text{thione}})-\text{S})$	496	509	490	531
$\nu(\text{C}=\text{S}) + \nu(\text{S}_{\text{thiole}}-\text{C}_{\text{thione}}-\text{S}_{\text{thiole}})$	993	1063	950	1039
$\nu(\text{Ni-S}_{\text{thiolate}})$	418	438	414	455
$\nu(\text{Ni-S}_{\text{thiolate}})$	317	337	318	310

The calculated peaks associated with the C-H stretch modes of the $[\text{CH}_3]$ groups were found at $3073/3098/3001 \text{ cm}^{-1}$ for B3LYP, CAM-B3LYP and PBE+U respectively. The bands at $1331/1374/1297 \text{ cm}^{-1}$ and $993/1063/950 \text{ cm}^{-1}$ were assigned to C=C and C=S stretching modes, respectively, and were compatible with other published results [40]. Valade et al. [41] also reported the C=C peak at 1430 cm^{-1} , and listed two peaks at 455 cm^{-1} and 310 cm^{-1} ; both were assigned as Ni-S vibration. The bands at $496/509/490 \text{ cm}^{-1}$ are characteristic of the fundamental vibrations of the thiocarbonate group $(-\text{S}-(\text{C}=\text{S}_{\text{thione}})-\text{S}-)$. The C=S stretching vibration is the characteristic vibration in the IR spectra of DMIT complexes. According to various

authors, several peaks appear in the 1050-995 cm^{-1} range, making it difficult to assign as C=S [42,43]. The results show the same tendency for different levels of calculations in describing the vibrational properties.

Considering now the electronic properties of metal complexes we first start analyzing the *HOMO* and *LUMO* Kohn–Sham frontier orbitals of $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$. As can be seen in Fig. 3, the *HOMO* is of π -character and mainly corresponds to the $\text{C}_2\text{S}_2^{2-}$ unit of dmit ligand and Ni(II) *d* orbital center. The *LUMO* is mainly contributed from the thiole ring of dmit ligand without metal participation. The same profile was observed by Fan et al. [45].

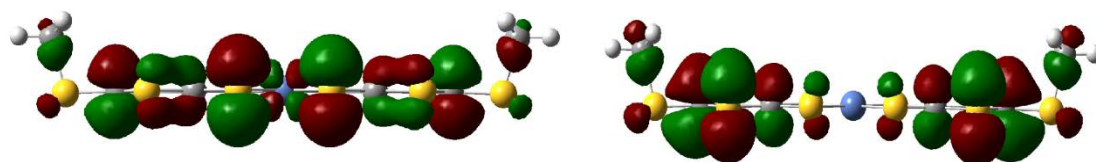


Fig. 3: Kohn–Sham frontier orbitals of isolated $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$ *HOMO* (left); *LUMO* (right).

Absorption electronic spectra were obtained from TD-DFT calculations performed at the optimized ground-state geometries (Fig. 4). Calculated oscillator strengths, transition energies, and wavefunction for the most relevant transitions of electronic absorption bands are listed in Table 3. The influence of the solvent environment on the absorption spectra was not considered in our calculations. Only transitions with significant oscillator strengths are presented. Fifty singlet electronic excited states were included.

Table 3: Excitation energy (E in eV), oscillator strength (f) and main configurations of the wavefunction of $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$ at TDDFT/B3LYP and TDDFT/CAM-B3LYP levels of theory (H=HOMO, L=LUMO).

B3LYP			CAM-B3LYP		
E	f	Main configurations	E	f	Main configurations
1.51	0.33	H→L	2.03	0.50	H→L/H-1→L+1
2.49	0.12	H-1→L+1	3.35	0.07	H→L/H-1→L+1
3.29	0.08	H-6→L	4.69	0.15	H-7→L
3.97	0.04	H→L+5/H→L+8	4.79	0.54	H-5→L+2
4.32	0.60	H-5→L+2/H→L+8/H→L+10			

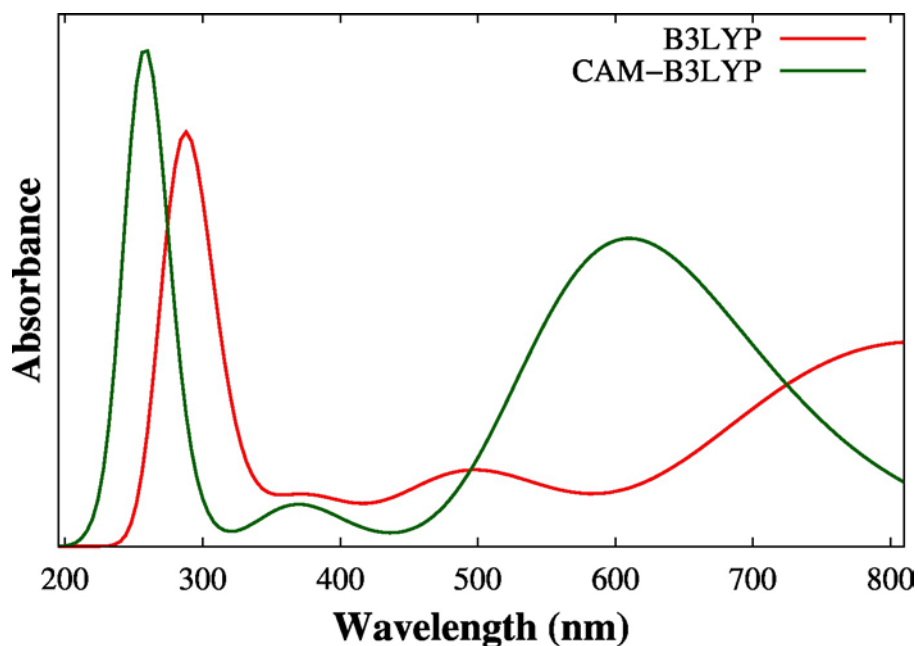


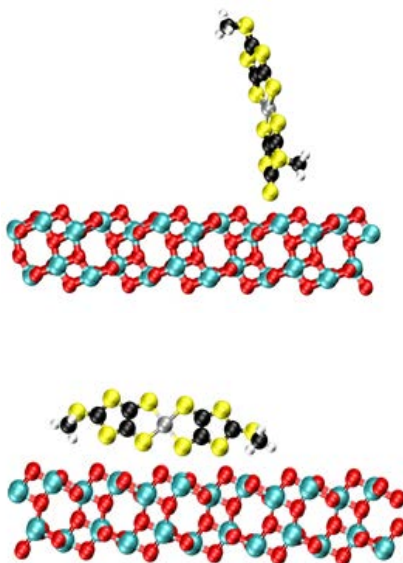
Fig 4.: Gas phase UV-vis absorption spectrum of $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$ computed at the TDDFT/B3LYP and TDDFT/CAM-B3LYP levels of theory.

As shown both in Table 3 and Fig. 4, the theoretical description of the UV-vis absorption spectrum of $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$ complex is, quantitatively, quite different for the two DFT functionals tested. The first absorption appears at a wavelength of ~ 800 nm when the B3LYP functional is used, However, the CAM-B3LYP functional offers a different picture, with a first, quite intense band appearing at ~ 590 nm, in much

good agreement with the experiment [46, 47]. Both functionals assign this excitation to a *HOMO-LUMO* $\pi \rightarrow \pi^*$ transition, but with significant contribution of the HOMO-1 \rightarrow LUMO+1 transition in the case of the CAM-B3LYP functional. At shorter wavelengths, the general features of the CAM-B3LYP computed spectrum also agrees in a better way with the experimental data. There are some excitations that give rise to an absorption band around 380 nm assigned to transitions involving molecular orbitals arising from the interaction between *d* metal orbitals with *p* orbitals of dmit moieties [48]. However, in the experimental spectrum the band is found at 400 nm and shows a higher intensity than the 600 nm band, pointing to some defects still present in the model. In the high region energy, however, there is a much better agreement with the experiment, noting a prominent band at 4.79 eV (259 nm) that involves intra- and inter-ligand $\pi \rightarrow \pi^*$ excited determinants.

Adsorption energies, structural and electronic properties of [CH₃]₂[Ni(dmit)₂] supported on the TiO₂ (101) surface.

Three different types of adsorptions onto TiO₂ surface configurations were considered: linked by S_{thione} (in this case the methyl group has been displaced to a nearby S_{thiole} site to allow S_{thione} to interact with the TiO₂ surface), linked by S_{thiole}-S_{thiolate}, and planar. The optimized geometry of these structures are shown in Fig. 5 and the values of the adsorption energies are summarized in Table 4.



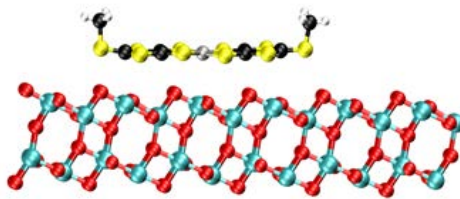


Fig. 5: Side-view optimized geometries for $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$ adsorbed on TiO_2 (101) anatase surface.

Table 4: Calculated adsorption energies (E_{ads} in eV) of $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$ on TiO_2 (101) surface.

Form of adsorption	E_{ads}
S_{thione}	0.59
$S_{\text{thiole}}-S_{\text{thiolate}}$	-0.92
Planar	-1.12

Table 4 shows that adsorption modes $S_{\text{thiole}}-S_{\text{thiolate}}$ and planar of Ni complex are moderately exothermic while binding through the S_{thione} atom is endothermic. Two main ingredients contribute to this different behavior. First, the stability of a complex with a bidentate ligand is invariably greater than the corresponding complex of monodentate (S_{thione}) ligands. Second, there is a large contribution from the dispersion forces for the $S_{\text{thiole}}-S_{\text{thiolate}}$ and planar modes, which is practically absent in the S_{thione} mode. In fact, the more efficient Van der Waals interaction in the case of planar adsorption makes it the preferred mode. There are no significant variations on the internal geometry of the Ni complex after adsorption on TiO_2 (101) surface.

To analyze the electronic structure, density of states (DOS) were calculated for the optimized structure of clean TiO_2 (101) surface, isolated $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$ complex, and $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$ complex adsorbed on TiO_2 (101) surface. The corresponding DOS plots for the isolated $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$ complex and the TiO_2 surface are reported in Fig. 6. The DOS of clean TiO_2 (101) shows a broad, filled valence band and a broad, empty conduction band, separated by a gap of 2.48 eV, which is consistent with previous theoretical works [49]. Inspection of Fig. 6 shows that the *HOMO* of the $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$ complex appears ~ 0.98 eV above the TiO_2 valence

band edge, lying within the band gap. In its turn, the *LUMO* level lies just below the conduction band edge.

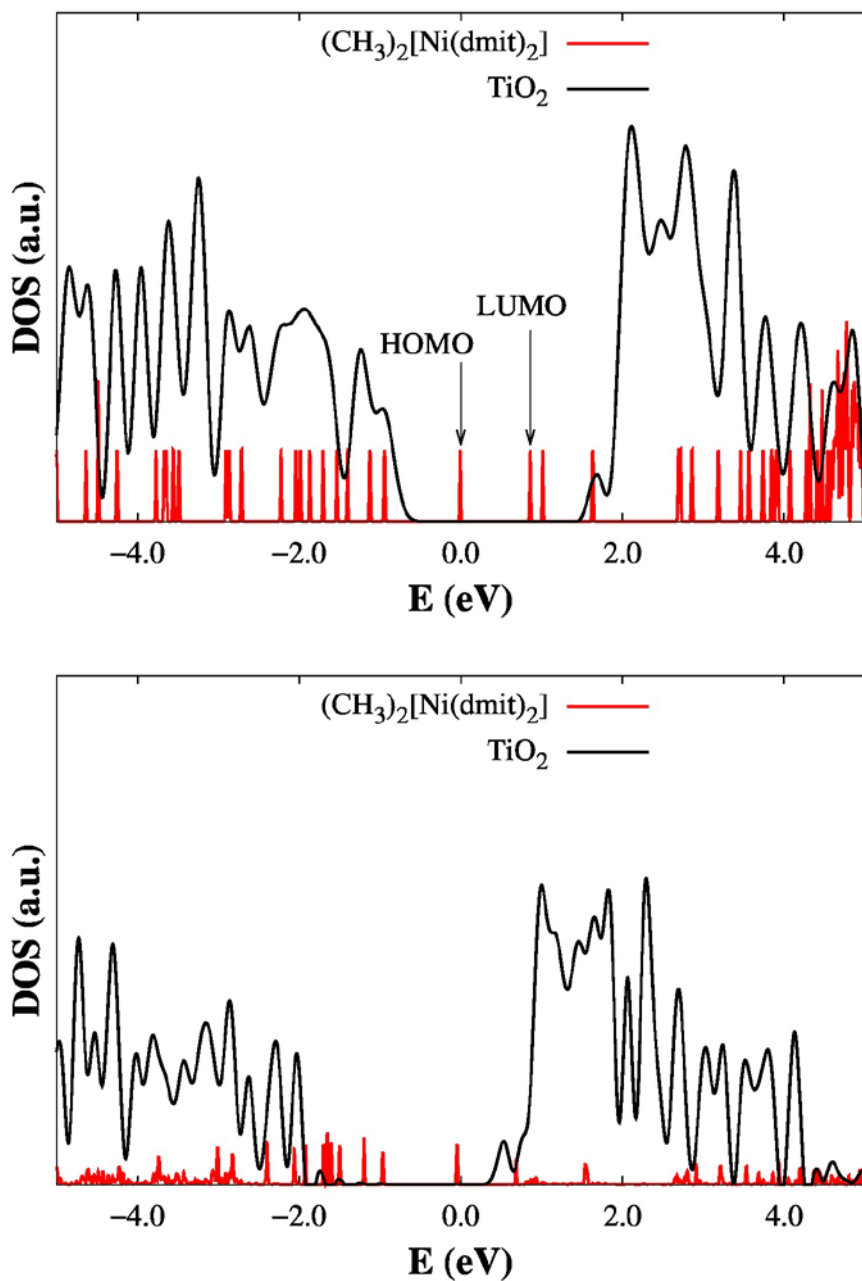


Fig. 6: Top: calculated DOS for the isolated $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$ complex and clean TiO_2 (101) surface; bottom: after adsorption. The zero of the energy is set at the Fermi level of the complex.

The performance of DSSC is associated with alignment of the metal complex *HOMO-LUMO* levels with respect to the TiO_2 band edges; this property is the key for

an efficient electron injection. The *LUMO* of metal complex has to be higher in energy than the conduction band edge of the semiconductor so the electron transfer occurs between *LUMO* and the TiO₂ conduction band. Therefore, charge injection times depend upon the electronic overlap between the TiO₂ and dye orbital. The analysis of DOS shows that the metal complex *LUMO* position relative the TiO₂ (101) surface is energetically unfavorable for electron injection on TiO₂ conduction band. After adsorption, the DOS, bottom of Fig 6, is found to be quite similar to that for the isolated components. The HOMO-LUMO gap of nickel complex decreases slightly while the LUMO level is somewhat pushed into the conduction band of the titania surface. As can be seen, some new occupied states appear at the TiO₂ band gap after adsorption of [CH₃]₂[Ni(dmit)₂] complex. These bands are contributed almost exclusively from the Ni complex orbitals. We also observe that metal complexes have a strong overlap with the conduction band. On the other hand, it is worth noting that the presence of [CH₃]₂[Ni(dmit)₂] affects the magnitude of the band gap. In any case, the calculated DOS for the different modes of adsorption are quite similar.

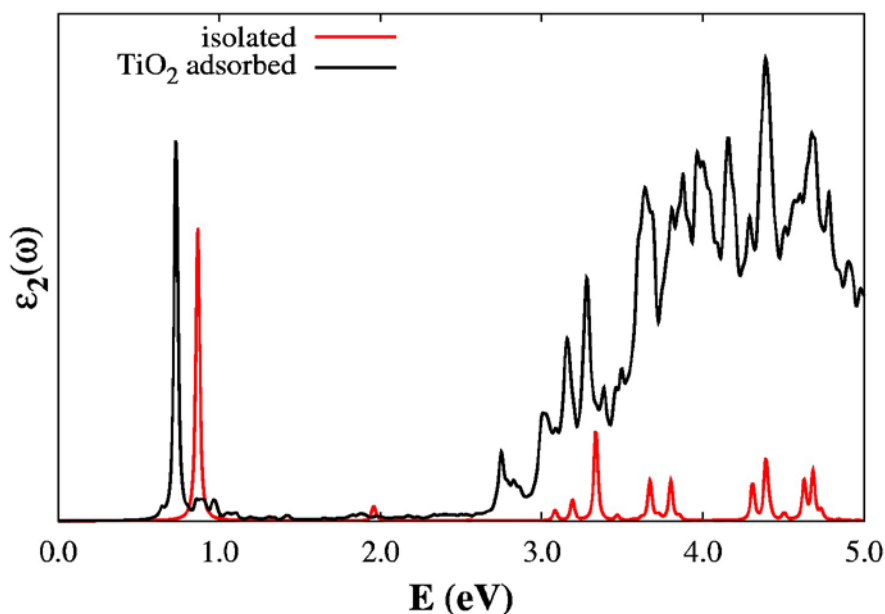


Fig. 7: Optical absorption spectra of isolated [CH₃]₂[Ni(dmit)₂] complex and adsorbed onto the anatase TiO₂ (101) surface, planar mode, computed at the PBE+*U* level of theory.

Finally, the UV–vis absorption spectra for the [CH₃]₂[Ni(dmit)₂] nickel complex adsorbed on the TiO₂(101) anatase surface in the planar configuration was

obtained. The imaginary part of the dielectric function, $\epsilon_2(\omega)$, computed from PBE+ U calculations is reported in Fig. 7. In order to make the analysis easier, the optical spectrum of the isolated nickel complex was also computed and plotted in Fig. 7.

We start with the spectrum of the isolated nickel complex. The lowest energy peak appears at 0.87 eV that according to the DOS analysis is assigned to a *HOMO-LUMO* transition. This is followed by a small peak at 1.95 eV, and several absorptions at energies higher than 3 eV. If we compare this spectrum with that obtained from TDDFT B3LYP calculations (Fig. 4), a general lowering of the transition energies is observed. Beyond the specified technical differences in the setup calculations: plane-waves *vs.* contracted Gaussian basis sets, and theoretical framework, the main responsible for this lowering is the exchange-correlation functional used in the calculations [50]. While hybrid exchange-correlation functionals use to overestimate HOMO-LUMO gaps, plane GGA functionals clearly underestimate them, which leads to a shift in the spectra [51,52]. Notice that although a Hubbard type correction is present in the PBE calculations, the U term is only applied to Ni atoms, while the lowest transitions mainly are of $\pi \rightarrow \pi^*$ character, primordially involving carbon and sulfur ligand atoms.

After deposition of the nickel complex on the $\text{TiO}_2(101)$ surface the optical spectrum shows a low energy strong absorption at 0.73 eV that corresponds to the HOMO-LUMO excitation of the free complex. Below 3 eV, the observed peaks are quite similar to those of the free complex but at lower energies, indicating that the interaction with the titanium oxide surface induces a general red shift of the electronic bands in agreement with the behavior reported for CdSe clusters supported on $\text{TiO}_2(110)$ rutile surface [51]. At energies higher than 3 eV the spectrum show complex features resulting from the superposition of the electronic absorption of both the nickel complex and the $\text{TiO}_2(101)$ surface [51]. This assignment is consistent with the DOS plots computed for $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$ nickel complex adsorbed on the $\text{TiO}_2(101)$ anatase surface reported at the bottom of Fig. 6. In any case, it is worth to remark the absence of significant electronic absorptions in the visible spectrum window (1.6-3.1 eV), indicating the unsuitability of this type of nickel complex as DSSC sensitizer. This situation, however, may be different if the complex is chemically bound to the TiO_2 surface, a possibility that will be the subject of a future work.

Conclusions

Density functional calculations have been carried out to explore the structural, electronic, and optical properties of $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$ complex, both isolated and supported on $\text{TiO}_2(101)$ anatase surfaces. With this aim, two different theoretical frameworks have been undertaken: molecular B3LYP calculations with a contracted Gaussian type basis set and periodic PBE+ U slab calculations using plane-waves. From a structural point of view both approaches give geometrical parameters and vibrational frequencies in agreement with experimental data. TDDFT/B3LYP calculations of $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$ isolated complex shows an intense and low energy absorption at 1.51 eV mainly arising from a *HOMO-LUMO* $\pi \rightarrow \pi^*$ excitation. The *HOMO* isosurface clearly shows the electron density being localized on $\text{C}_2\text{S}_2^{2-}$ moieties of dmit ligand and Ni (II) d orbital. In contrast, the *LUMO* is mainly composed by the dmit ligand thiole ring without nickel contribution.

For the deposition of $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$ on the $\text{TiO}_2(101)$ surface three different adsorption configurations were considered. Our calculations indicate that the most stable adsorption mode corresponds to a planar form where the dmit π system interacts with the TiO_2 surface, mainly through long-range dispersion forces. Partial DOS plots show that the LUMO of $[\text{CH}_3]_2[\text{Ni}(\text{dmit})_2]$ lies at the conduction band edge of TiO_2 which is unfavorable for a hypothetical electron injection onto the semiconductor surface. On the other hand, the HOMO of the complex also falls within the TiO_2 band gap, making it that the lowest (and most intense) electronic absorption of the nickel complex supported on the TiO_2 is similar to that of the unsupported complex. The effect of the $\text{TiO}_2(101)$ surface on the absorption spectra of the nickel complex is practically limited to a red shift of about 0.1-0.3 eV. The absence of significant electronic absorption in the visible spectrum region allows us to discard this complex as sensitizer in DSSC.

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