


Tuning Activity and Selectivity during Alkyne Activation by Gold(I)/Platinum(0) Frustrated Lewis Pairs

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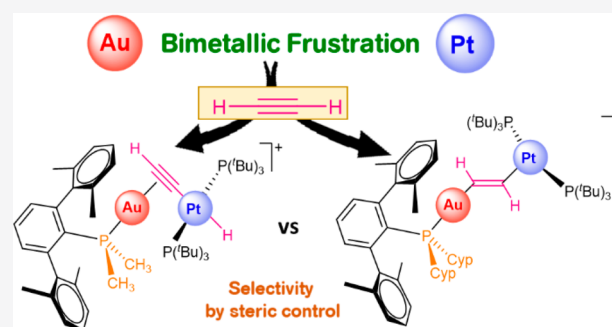
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ABSTRACT: Introducing transition metals into frustrated Lewis pair systems has attracted considerable attention in recent years. Here we report a selection of three metal-only frustrated systems based on Au(I)/Pt(0) combinations and their reactivity toward alkynes. We have inspected the activation of acetylene and phenylacetylene. The gold(I) fragments are stabilized by three bulky phosphines bearing terphenyl groups. We have observed that subtle modifications on the substituents of these ligands proved critical in controlling the regioselectivity of acetylene activation and the product distribution resulting from C(sp)–H cleavage of phenylacetylene. A mechanistic picture based on experimental observations and computational analysis is provided. As a result of the cooperative action of the two metals of the frustrated pairs, several uncommon heterobimetallic structures have been characterized.



INTRODUCTION

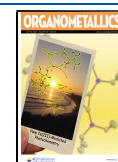
Frustrated Lewis pairs (FLPs) emerged more than a decade ago as a paradigmatic example of chemical cooperativity, permitting bond activation and catalysis in the absence of transition metals.¹ However, the incorporation of the latter as core components of FLP systems has also attracted considerable attention in the last few years.^{2,3} In fact, the concept of frustration may be applied to a wide variety of transition-metal-mediated transformations that were described considerably earlier than the emergence of FLP chemistry and have so far been associated with the broader concept of chemical cooperativity. From an awareness of this sometimes diffuse frontier between “traditional” transition-metal cooperative chemistry and metallic frustration, the introduction of the latter concept remains useful as a driving force for the discovery of novel transformations. That being said, introducing transition metals into frustrated designs largely increases the amount and structural diversity of Lewis acid/base combinations available. In addition, it provides an array of elementary reactions accessible for transition metals that is foreseen to extend the catalytic usefulness of main-group FLPs beyond their current status. However, a fundamental knowledge on transition-metal FLPs regarding mechanistic aspects, solution dynamics, acid–base interactions, and selectivity effects is rather underexplored in comparison to main-group FLPs, despite the fact that this information is vital for expeditious catalyst development.

In this regard, the Wass group demonstrated that subtle ligand modifications have a strong effect on the ability of zirconocene-based FLPs toward dihydrogen splitting.⁴ This type of system

was also examined to clarify the nature of Lewis acid–base interactions by DOSY NMR spectroscopy.⁵ Going beyond monometallic systems, our group focused on gaining a fundamental knowledge of FLPs in which the two Lewis components are based on transition metals. Thus, we recently reported the first example of its kind by combining Au(I) and Pt(0) species as the acidic and basic sites, respectively.⁶ To achieve frustration, we targeted sterically hindered phosphine ligands for both gold and platinum monometallic complexes. Our experimental/computational investigations regarding the heterolytic splitting of dihydrogen mediated by these pairs led us to propose a genuine bimetallic FLP-type pathway⁷ analogous to the models assumed for main-group counterparts.⁸ Moreover, we could analyze the strong influence that Au···Pt interactions have on the activation capacity of the bimetallic pairs, as well as the solution dynamic equilibria between the metal-only Lewis pairs and the individual monometallic fragments. This is a particularly important aspect in the field of FLPs that has been widely investigated for metal-free systems, where the term “thermally induced FLPs”⁹ was coined to refer those pairs in which the Lewis adduct is the resting state. Despite this fact, many of these pairs exhibit a rich FLP reactivity¹⁰ and in some

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cases catalytic performance superior to that of their fully frustrated counterparts.¹¹

In this study, we extend our fundamental knowledge on transition-metal-only FLPs (TMOFLPs) by exploring regioselectivity effects derived from ligand modification during the activation of alkynes, also model substrates widely investigated in the field of frustrated systems.^{1,12} We have focused on the effects derived from varying the degree of frustration, for which we have used terphenylphosphine ligands $\text{PMe}_2\text{Ar}^{\text{Xyl}2}$ (**a**), $\text{PMe}_2\text{Ar}^{\text{Dipp}2}$ (**b**), and $\text{PCyp}_2\text{Ar}^{\text{Xyl}2}$ (**c**) (Cyp = cyclopentyl), with different steric profiles, to stabilize electrophilic gold fragments **1** (Figure 1). These acidic complexes combined with the basic

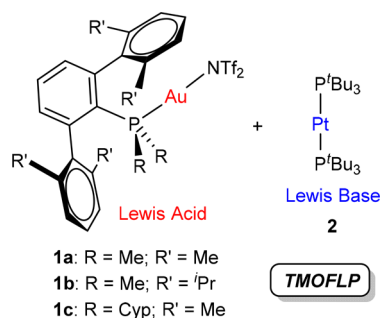


Figure 1. Au(I)/Pt(0) transition-metal-only frustrated Lewis pair (TMOFLP) combinations used in this work (Cyp = cyclopentyl; $\text{NTf}_2^- = \text{N}(\text{SO}_2\text{CF}_3)_2^- = \text{triflimide}$).

Pt(0) compound $[\text{Pt}(\text{P}^t\text{Bu}_3)_2]$ (**2**) promote the cooperative activation of terminal alkynes. Our studies demonstrate the potential of using transition-metal Lewis acids (i.e., $[\text{PR}_3\text{Au}]^+$ fragments) to control the selectivity in the activation of small molecules by tuning the steric properties of the ancillary ligands. This is particularly appealing in view of the challenging available protocols to synthesize the typically used acidic boranes and their scant stability toward moisture.¹³

RESULTS AND DISCUSSION

To investigate regioselectivity effects during alkyne activation, we first examined the reactivity of the three Au/Pt bimetallic pairs depicted in Figure 1 toward acetylene, a reaction that we had previously reported with the **1b**:**2** combination.^{6a} When a dichloromethane or benzene solution of the latter pair is exposed to acetylene (0.5 bar, 25 °C), a rapid color change from bright yellow to intense orange takes place. Multinuclear NMR spectroscopic analysis revealed the formation of a clean mixture of two structurally different isomers, namely a bridging σ,π -acetylide (**3b**) and a rather unusual heterobimetallic vinylene ($-\text{CH}=\text{CH}-$) (**4b**), which are produced in a 4:1 ratio

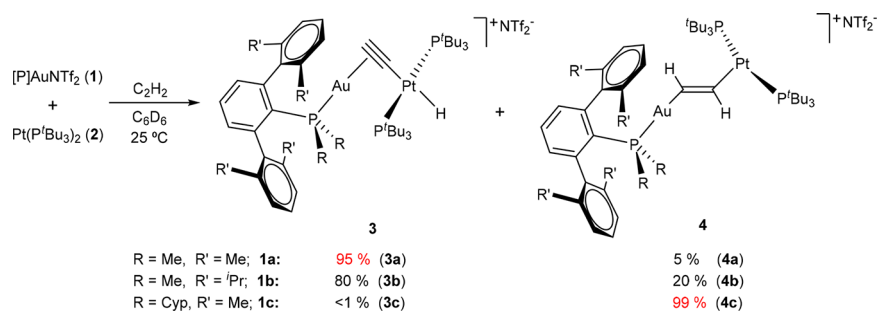
(Scheme 1). These metallic species are highly reminiscent of the organic products derived from the reactivity of traditional phosphine/borane FLPs with alkynes, where the prevalence of one or the other isomer typically depends on the basicity of the phosphine.¹⁴

We now tested the analogous reactivity using the gold precursors $[(\text{PMe}_2\text{Ar}^{\text{Xyl}2})\text{Au}(\text{NTf}_2)]$ (**1a**) and $[(\text{PCyp}_2\text{Ar}^{\text{Xyl}2})\text{Au}(\text{NTf}_2)]$ (**1c**) in our search for regioselectivity effects, while the basicity of the metallic base was kept unaltered (**2**). Moreover, the acidities of the gold precursors **1a–c** barely differ from one another;¹⁵ thus, any anticipated outcomes mostly build on steric grounds. In fact, we found a drastic change in product distribution from the less hindered system (**1a**, $\text{PMe}_2\text{Ar}^{\text{Xyl}2}$) to the more congested one (**1c**, $\text{PCyp}_2\text{Ar}^{\text{Xyl}2}$), as determined by NMR spectroscopy. While the former yields around 95% of the bridging Au/Pt acetylide **3a** and only a residual amount of the vinylene (**4a**, <5%), the more hindered pair comprising the $[(\text{PCyp}_2\text{Ar}^{\text{Xyl}2})\text{Au}]^+$ fragment (**1c**) fully reversed the selectivity toward the exclusive formation of the corresponding vinylene **4c** (Scheme 1). Attempts to isolate **3c** by the reaction of independently prepared compounds $[(\text{PCyp}_2\text{Ar}^{\text{Xyl}2})\text{Au}(\text{C}\equiv\text{CH})]$ (**5c**) and $[\text{Pt}(\text{P}^t\text{Bu}_3)_2\text{H}][\text{NTf}_2]$ ¹⁶ (**6**) proved unsuccessful and resulted in intractable mixtures.

As mentioned above, this dramatic shift in regioselectivity seems to be dominated by steric effects, which contrasts with prior strategies to modulate alkyne activation by FLPs that mostly rely on phosphine basicity. More importantly, it evinces the potential of FLP systems that incorporate transition-metal Lewis acids to easily tune the selectivity during bond activation processes and, as such, in subsequent catalytic applications that incorporate those activation events. As was pointed out earlier, this could be seen as a key advantage in comparison to traditional FLP designs that usually involve fluorinated boranes, since accessing these moieties already entails substantial synthetic challenges and limitations, not to mention their limited stability toward moisture and air.¹³ In stark contrast, the preparation of terphenylphosphines PR_2Ar^r is straightforward and highly versatile,¹⁷ while the resulting gold precursors **1** are readily obtained in high yields and exhibit stability toward water or under moderate oxidizing conditions.¹⁸

The nature of the new heterobimetallic compounds **3a** and **4c** was ascertained by a comparison of their ^1H and $^{31}\text{P}\{^1\text{H}\}$ NMR signals with those derived from their analogous species based on $\text{PMe}_2\text{Ar}^{\text{Dipp}2}$: that is, **3b** and **4b**, respectively.^{6a} The heterobimetallic nature of compounds **4** is evinced by the ^{195}Pt satellites that flank the $^{31}\text{P}\{^1\text{H}\}$ resonances associated with terphenylphosphines, which appear at 2.1 (**4b**, $^4J_{\text{Ppt}} = 282$ Hz) and 51.7 (**4c**, $^4J_{\text{Ppt}} = 277$ Hz) ppm. The bridging vinylene ($-\text{CH}=\text{CH}-$) moiety displays a distinctive pair of ^1H NMR

Scheme 1. Regioselectivity in the Activation of Acetylene by TMOFLPs **1**:**2**



signals in the region between 4.0 and 4.5 ppm that reveal scalar coupling to the ^{195}Pt center in the range 120–200 Hz (see the [Experimental Section](#) for details). By analogy, we attribute a $^{31}\text{P}\{^1\text{H}\}$ NMR resonance at 3.67 ppm to the minor species (ca. 5%) in the $\text{PMe}_2\text{Ar}^{\text{Xyl}2}$ system (**4a**), with an identical ^1H NMR pattern comprised of signals at 4.54 and 4.37 ppm, though their corresponding ^{195}Pt satellites could not be observed due to the low concentration of isomer **4a**. Corresponding ^{13}C NMR resonances for compounds **4b,c** emerge at ca. 155 ($^1J_{\text{CH}} \approx 175$ Hz) and 115 ($^1J_{\text{CH}} \approx 190$ Hz) ppm, respectively, supporting the proposed formulation and the sp^2 hybridization of the carbon atoms.

The molecular structure of compounds **3a** and **4c** was further corroborated by X-ray diffraction studies ([Figure 2](#)). The

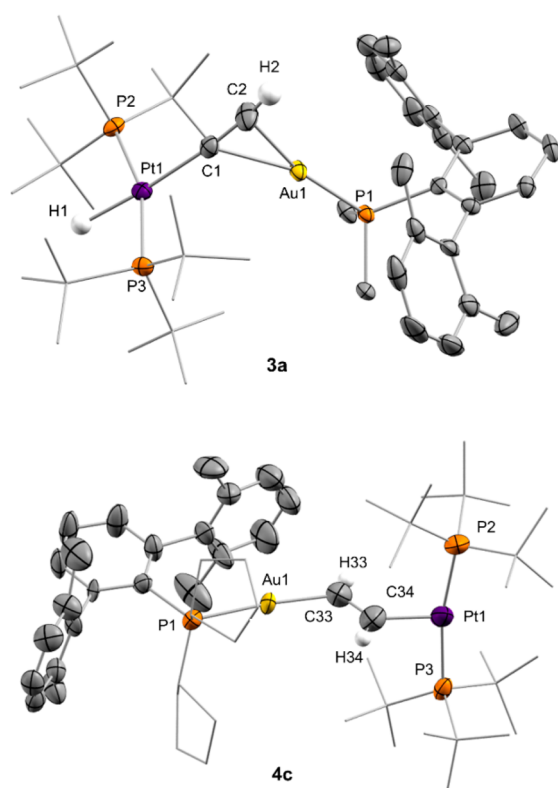


Figure 2. ORTEP diagrams of compounds **3a** and **4c**. For the sake of clarity most hydrogen atoms and triflimide anions are excluded and some substituents have been represented in wireframe format, while thermal ellipsoids are set at 50% probability.

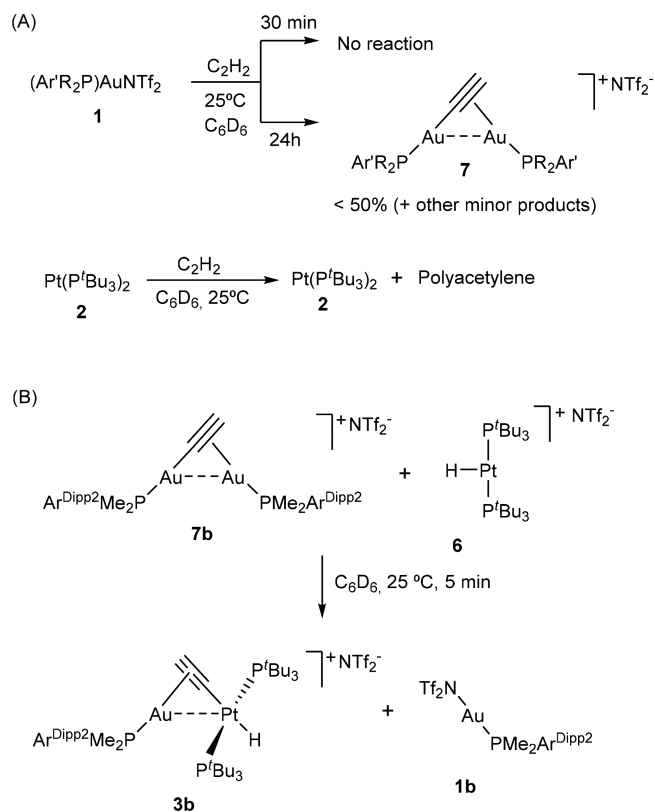
presence of the σ,π -acetylide or vinylene linker distorts the linearity around the platinum center: P–Pt–P angles of around 165° , shifted from the ideal 180° due to the steric pressure exerted by the bulky gold fragments. The Pt1–C1 (2.016(6) Å) and Au1–C1 (2.311(5) Å) bond distances in **3a** appear slightly shortened in comparison to **3b** ($d_{\text{Pt1}-\text{C1}} = 2.044(7)$ Å; $d_{\text{Au1}-\text{C1}} = 2.360(7)$ Å). The average C=C bond distance of the vinylene linkers in the two crystallographically independent molecules of **4c** accounts for 1.278(13) Å, comparable to those of **4b** (1.287(11) Å). These and other geometric features are similar to those of previously reported related species.¹⁹

We had previously observed that the three investigated terphenylphosphines permit control of the equilibrium between complete frustration and bimetallic adduct formation.⁷ Thus, while a dative Pt→Au bond is immediately formed between $[(\text{PMe}_2\text{Ar}^{\text{Xyl}2})\text{Au}(\text{NTf}_2)]$ (**1a**) and $[\text{Pt}(\text{P}^t\text{Bu}_3)_2]$ (**2**), the

formation of an identical adduct based on $\text{PCyp}_2\text{Ar}^{\text{Xyl}2}$ is endergonic and could not be experimentally detected. An intermediate situation is reached for the medium-sized phosphine $\text{PMe}_2\text{Ar}^{\text{Dipp}2}$, where the prevalence of the mono-metallic fragments or the bimetallic adduct depends upon experimental conditions. In this context, we have observed that TMOFLPs (**1:2**) based on the gold precursors $[(\text{PMe}_2\text{Ar}^{\text{Dipp}2})\text{Au}(\text{NTf}_2)]$ (**1b**) and $[(\text{PCyp}_2\text{Ar}^{\text{Xyl}2})\text{Au}(\text{NTf}_2)]$ (**1c**) are considerably more active toward alkyne activation in comparison to that one built on $[(\text{PMe}_2\text{Ar}^{\text{Xyl}2})\text{Au}(\text{NTf}_2)]$ (**1a**). While full conversion toward compounds **3** and **4** was recorded by the time of placing the sample in the NMR probe (<5 min) in the case of using **1b:2** or **1c:2**, the analogous transformation essayed with **1a** required up to 24 h to reach completion under otherwise identical conditions (C_2H_2 , 0.5 bar, 25°C , toluene or C_6D_6). This fact speaks in favor of a genuine FLP mechanism that imposes an energetic demand to overcome the Pt→Au bond cleavage prior to acetylene activation, a requirement that only applies to the less hindered gold precursor **1a**. Similar effects have been observed in the gold-catalyzed addition of P-nucleophiles to alkynes.²⁰

In addition, it is key to highlight that the cooperative reactivity depicted in [Scheme 1](#) contrasts with that of the individual Au(I) or Pt(0) fragments ([Scheme 2](#)). For instance, $[\text{Pt}(\text{P}^t\text{Bu}_3)_2]$ (**2**) readily catalyzes acetylene polymerization, evinced by the rapid formation of a purple-black solid accompanied by the disappearance of a ^1H NMR resonance at 1.34 ppm due to C_2H_2 , while signals due to **2** remained unchanged. At variance, no indication of polyacetylene formation is apparent when gold

Scheme 2. (A) Reactivity of Individual Gold (1) and Platinum (2) Compounds toward Acetylene and (B) Reaction of σ,π -Acetylide **7b and T-shaped Platinum Hydride **6****



is also present in solution. In the case of the individual gold compounds **1** there is no sign of chemical transformation in the short term (ca. 30 min), while at longer reaction times the gold triflimide precursors evolve to bridging σ,π -acetylide compounds **7** (Scheme 2A),²¹ albeit only in moderate yields accompanied by other unidentified gold-containing species.

Having the previous experimental findings on hand, we were interested in further understanding the cooperative action of TMOFLPs **1**:**2** to gain a fundamental knowledge of significance for future catalytic applications. As such, we initially wondered about the possible role of acetylide-bridged digold compounds such as **7** as precursors toward complexes **3** and **4**. In fact, the reaction between **7b** and [Pt(P^tBu₃)₂H][NTf₂] (**6**) immediately yielded the corresponding heterobimetallic σ,π -acetylide compound **3b**, where the unsaturated linker is now σ -bonded to the platinum center instead of the gold nucleus. Nevertheless, the complete absence of the vinylene isomer **4b** during the latter reaction would require an additional competing route to provide access to this unusual bimetallic motif, which actually is the exclusive isolated isomer for the bulkier PCyp₂Ar^{Xyl2}-based system (Scheme 1). Moreover, the formation of compounds **7** requires several hours to proceed to appreciable conversions, while the activation of acetylene by **1**:**2** pairs is immediate (<5 min), except for the gold precursor **1a** bearing PMe₂Ar^{Xyl2}, which as noted earlier takes around 24 h to convert into **3a** in the presence of [Pt(P^tBu₃)₂] (**2**).

To further investigate the mechanism and the reasons for the drastically different regioselectivity observed, we carried out computational studies (DFT, ω B97XD/6-31G(d,p)+SDD). Focusing on the system based on PMe₂Ar^{Dipp2}, we began searching for initial acetylene activation steps by approaching the acetylene molecule to the individual Au(I) (**1b**) and Pt(0) (**2**) fragments, assuming that Au–Pt dissociation is a prerequisite for alkyne activation, as deduced from the reduced reactivity of the pair **1a**:**2** in comparison to **1b**:**2** and **1c**:**2** and previous computational results.⁷ Interestingly, the formation of **3** and **4** seems to share a common intermediate, namely a Au(I) acetylene adduct of formula [(PR₂Ar')Au(C₂H₂)]⁺ (**9**). The formation of this type of π complex has been previously proposed in the context of alkyne^{12,14a} and alkene²² activation by P/B pairs, but no experimental proofs of their existence have been reported. In an attempt to spectroscopically identify such an intermediate, we recorded the formation of a new gold-containing species by low-temperature NMR (–80 °C) as the major species (ca. 80%) upon exposing a CD₂Cl₂ solution of **1b** to an acetylene atmosphere. This species exhibits a distinctive ¹H NMR signal at 3.43 ppm that correlates with a ³¹P NMR resonance at –0.3 ppm (in comparison to δ_p –9.9 ppm for **1b**) and presents dynamic exchange with free acetylene (δ_H 2.11 ppm) as seen by EXSY NMR experiments, while other ¹H NMR resonances are comparable to those of **1b**. This finding constitutes an additional benefit of TMFLPs for mechanistic investigations in frustrated systems, since they provide additional modes of stabilizing otherwise fleeting intermediates.

With the acetylene adduct **9b** as the starting point, our calculations indicate that the attack of the platinum compound **2** over **9b** leads to either of the two bimetallic isomers **3b** and **4b** depending on the trajectory followed by **2** while it approaches **9b** (Figure 3). Thus, if **2** approaches the acetylene adduct along the Au–C₂H₂ direction, the corresponding vinylene **4b** forms ($\Delta G^\ddagger = 23.4$ kcal mol^{–1} from the **1b**:**2** Lewis pair + acetylene), in a process somehow reminiscent of the gold-mediated nucleophilic attack over activated alkynes (e.g., gold-catalyzed

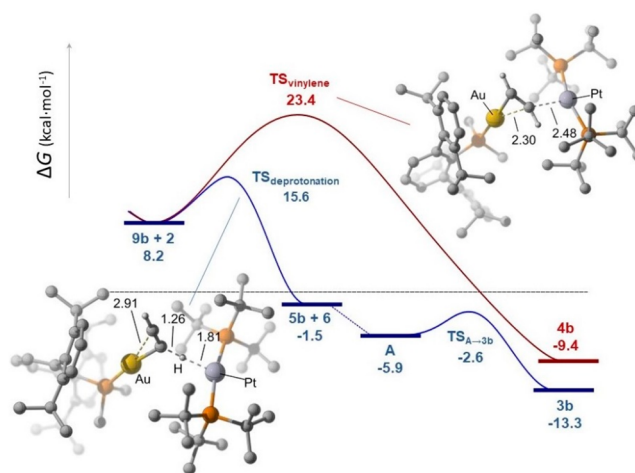
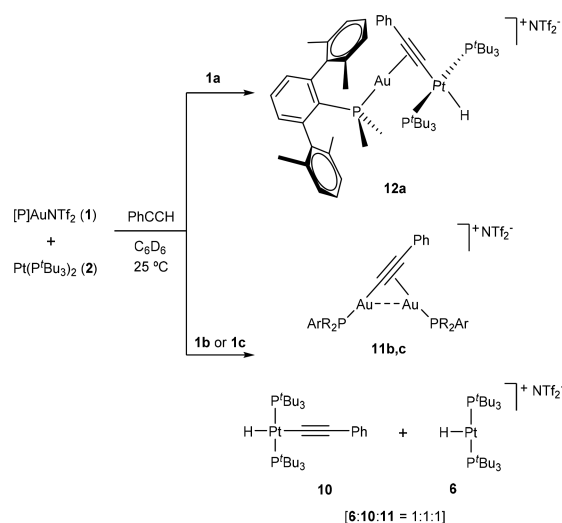


Figure 3. Energy profiles for the formation of the acetylide and vinylene complexes **3b** and **4b** from the common intermediate **9b** and platinum(0) compound **2**. The origin of energies is set to the **1b**:**2** Lewis pair + acetylene.

hydroamination).²³ In contrast, the alignment of the basic **2** in an orthogonal disposition with respect to the Au–C₂H₂ bond results in deprotonation of the activated acetylene (15.6 kcal mol^{–1} for the pair **1b**:**2**) to yield the corresponding Au(I) terminal acetylide [(PMe₂Ar^{Dipp2})Au(C≡CH)] (**5b**) and [Pt(P^tBu₃)₂H][NTf₂] (**6**). These two fragments readily rearrange to intermediates [(PR₂Ar')Au(μ-η¹:η²-C≡CH)Pt(H)(P^tBu₃)₂]⁺ (**A**) that subsequently evolve to compounds **3** by rapid σ,π isomerization of the bridging μ -C≡CR unit. In experimental agreement, the reaction of independently synthesized [(PMe₂Ar^{Dipp2})Au(C≡CH)] (**5b**) and [Pt(P^tBu₃)₂H][NTf₂] (**6**) rapidly yields complex **3b**. As mentioned above, this does not apply to compound **5c**, whose reaction with **6** resulted in a complex mixture of products that include decomposition into elemental gold. Nevertheless, the corresponding [(PCyp₂Ar^{Xyl2})Au(μ-η²:η¹-C≡CH)Pt(H)(P^tBu₃)₂]⁺ (**3c**) has not been detected during the bimetallic activation of acetylene. An alternative orthogonal mechanism involving the initial oxidative addition of acetylene over **2**, followed by cis–trans isomerization and coupling with **1**, was ruled out on the basis of significantly higher overall activation barriers (see Figure S2).

Information on the activation of the simplest alkyne (C₂H₂) by FLPs is rather scarce.^{12a} For the sake of completeness and to better compare our results with prior studies on main-group FLPs, we decided to test other more commonly employed triply bonded hydrocarbons. With regard to internal alkynes, all our attempts to access bimetallic vinylenes were unsuccessful. Reactions of **1**:**2** pairs with diphenylacetylene, 2-butyne, and 1,4-diphenylbutadiyne did not result in the formation of any new species even under more forcing experimental conditions. At variance, addition of phenylacetylene to an equimolar mixture of **1** and **2** provided phosphine-dependent divergent outcomes derived from C(sp)–H bond cleavage (Scheme 3). In parallel with acetylene activation with the more congested pairs based on **1b,c**, the reaction of these systems with PhC≡CH was also immediate. In contrast, while acetylene activation took up to 24 h for the nonfrustrated **1a**:**2** pair, the reaction was complete after around 15 min in the case of phenylacetylene.

The appearance of a distinctive low-frequency ¹H NMR resonance at around –10 ppm flanked by ¹⁹⁵Pt satellites (¹J_{HPt} = 608 Hz, **1a**:**2**; ¹J_{HPt} = 533 Hz, **1b**:**2** and **1c**:**2**) prompted us to

Scheme 3. Product Distribution from the Activation of Phenylacetylene by TMOFLP Pairs 1:2


believe that corresponding heterobimetallic σ,π -acetylide complexes were formed in all cases. However, a more careful analysis revealed unexpected differences in product distribution for the less bulky terphenylphosphine $\text{PMe}_2\text{Ar}^{\text{Xyl}2}$ in comparison to the more hindered systems (Scheme 3). Extracting the crude reaction mixtures with pentane permitted isolation of the same platinum-containing compound **10** for the pairs **1b:2** and **1c:2**, while the less hindered gold fragment **1a** did not lead to any metallic species soluble in nonpolar hydrocarbon solvents. An infrared band at 2090 cm^{-1} was recorded for a Pt-hydride ligand in **10**, while characteristic $^{13}\text{C}\{^1\text{H}\}$ NMR resonances at 118.8 and 117.6 ppm accounted for a σ -bonded acetylide ligand. On the basis of these spectroscopic features and the high solubility, we proposed a molecular formulation for **10** as $[\text{Pt}(\text{P}^t\text{Bu}_3)_2(\text{H})-(\text{C}\equiv\text{CPh})]$: that is, the formal oxidative addition of phenylacetylene over Pt(0) compound **2**. This assumption was further corroborated by X-ray diffraction analysis (see Figure S3). Nevertheless, it is important to remark that compound **2** does not react with phenylacetylene even after longer reaction times (48 h) or at elevated temperatures ($80\text{ }^\circ\text{C}$), which suggests a cooperative action between platinum and gold to account for alkyne C–H activation.

The only detectable gold-containing species in these reactions were assigned to the corresponding bridging σ,π -acetylide digold complexes **11**. Those were characterized by $^{31}\text{P}\{^1\text{H}\}$ NMR signals at 0.4 (**11b**) and 53.9 (**11c**), shifted to higher frequencies with respect to their precursors **1** (cf. -11.9 for **1b** and 48.8 ppm for **1c**), in agreement with other related examples.²⁴ Also similar to those results, the presence of a single ^{31}P resonance for each compound suggests rapid exchange of the σ,π coordination in solution.¹⁸ However, the process is frozen in the solid state. An ORTEP representation of the molecular structure of **11b** is shown in Figure 4a. The σ,π coordination of the acetylide is reflected by a nonsymmetric arrangement characterized by bond distances of 2.021(10) and 2.209(9) Å for the Au2–C1 and Au1–C1 bond distances, respectively. The Au1 center is also connected to C2 by a slightly longer bond length of 2.310(9) Å. The presence of an aurophilic interaction can be inferred from a Au1–Au2 distance of 3.366(1) Å, which is faintly elongated in comparison to its related acetylide analogue $[\text{Au}_2\{\mu\text{-C}\equiv\text{CH}\}]$ ($d_{\text{Au1-Au2}} = 3.31$ Å), likely as a result of the higher steric pressure exerted by

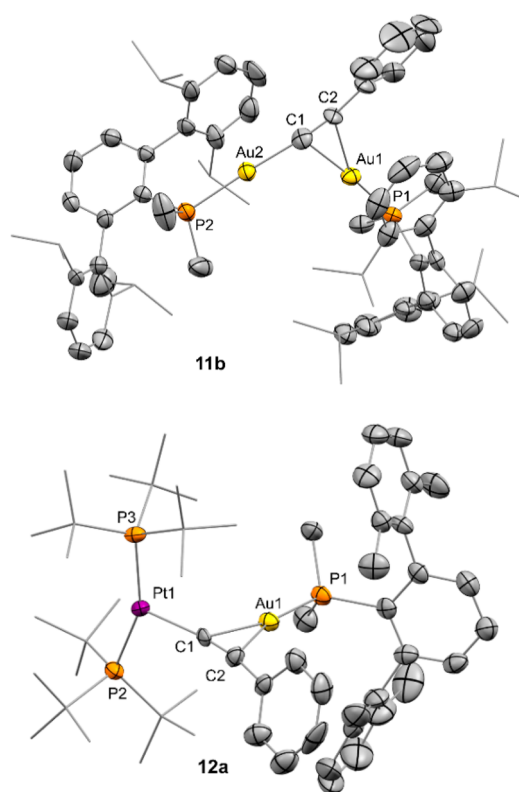


Figure 4. ORTEP diagrams of compounds **11b** and **12a**. For the sake of clarity hydrogen atoms, solvent molecules, and triflimide anions are excluded and some substituents have been represented in wireframe format, while thermal ellipsoids are set at 50% probability. The hydride ligand bound to platinum in **12a** could not be located in the Fourier electron density map.

phenylacetylide. Digold complexes **11b,c** were accompanied by equimolar amounts of $[\text{Pt}(\text{P}^t\text{Bu}_3)_2(\text{H})]^+$ (**6**) that could not be washed out with pentane, as evinced by a broad low-frequency ^1H NMR resonance at 1.16 ppm ($^1J_{\text{Pt}} = 6.5$ Hz).

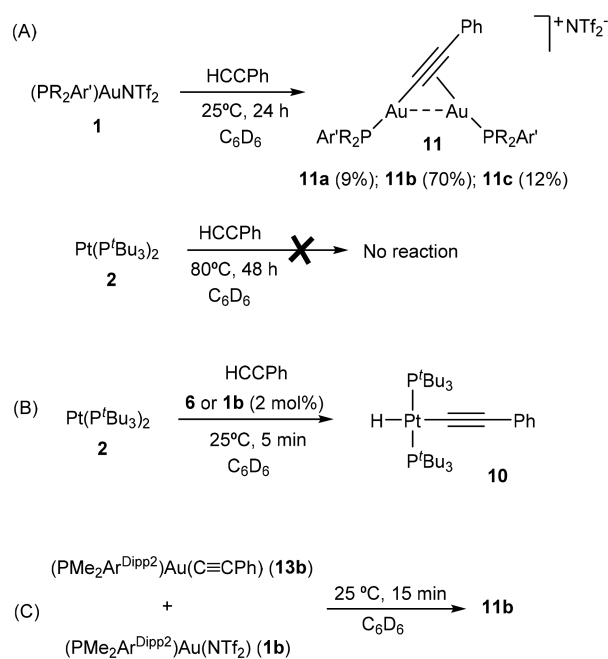
As noted earlier, the reaction of the less hindered **1a:2** pair with phenylacetylene yielded a divergent result. C(sp)–H activation became evident by the presence of a distinctive hydridic ^1H NMR resonance at -10.4 ppm ($^2J_{\text{HP}} = 14$ Hz, $^1J_{\text{HPt}} = 608$ Hz). However, a single platinum complex was formed in this case, which resonates at 82.2 ppm ($^1J_{\text{Pt}} = 2810$ Hz) in its $^{31}\text{P}\{^1\text{H}\}$ NMR spectrum and could not be washed out using nonpolar hydrocarbon solvents. The corresponding $\text{C}\equiv\text{C}$ stretching frequency rendered an infrared band shifted to lower wavenumbers ($\nu_{\text{C}\equiv\text{C}} 1982\text{ cm}^{-1}$; cf. 2048 cm^{-1} for **11b**, 2029 cm^{-1} for **11c**, and 2090 cm^{-1} for **10**), while sp-hybridized carbon atoms render $^{13}\text{C}\{^1\text{H}\}$ NMR resonances at lower frequencies (91.1 and 85.9 ppm) in comparison to compounds **11b,c** and **10** (116–125 ppm). These observations expose the divergent product distribution derived from using phosphines with different steric profiles. Since the recorded parameters equate with heterobimetallic σ,π -acetylide compounds **3**, we assumed an analogous structure for this complex (**12a**), a premise that we could substantiate by X-ray diffraction analysis (Figure 4b). The bulkier nature of the phenylacetylide moiety with respect to the unsubstituted acetylide in **3a** is likely the cause of a more intense distortion of the T-shaped platinum fragment, where a P–Pt–P angle of $156.71(8)^\circ$ is recorded (cf. $162.76(5)^\circ$ for **3a**). The close proximity between the acetylide phenyl fragment and the *ortho* substituents of one of the flanking

aryl rings of the terphenyl moiety may be responsible for the dissimilar product distributions found between $\text{PMe}_2\text{Ar}^{\text{Xyl}2}$ vs $\text{PMe}_2\text{Ar}^{\text{Dipp}2}/\text{PCyp}_2\text{Ar}^{\text{Xyl}2}$.

On the basis of our experimental and computational studies on acetylene activation, initial formation of a gold–alkyne adduct like **9** seems most plausible. The slightly higher acidity of phenylacetylene, as well as its higher size, may account for the prevalence of the deprotonation pathway in detriment of the 1,2-addition route toward vinylene structures, which were not detected. In turn, the dissimilar reaction products depicted in Scheme 3 might be understood on steric grounds. Thus, selective formation of compounds **11** for the bulkier $\text{PMe}_2\text{Ar}^{\text{Dipp}2}/\text{PCyp}_2\text{Ar}^{\text{Xyl}2}$ -based systems could be the result of a higher steric clash between the *tert*-butyl substituents on the platinum fragment and the terphenyl moiety of the gold-bound phosphine. In contrast, the reduced steric pressure introduced by $\text{PMe}_2\text{Ar}^{\text{Xyl}2}$ may permit easier access to the heterobimetallic compound **12a**, analogous to its unsubstituted acetylide version (**3a**).

We thought of interest to carry out several further experiments to shed some light into the operating cooperative mechanism. In the same manner as we observed for acetylene activation, the reactivity of the individual metallic fragments starkly contrasts with that of the bimetallic pairs. Accordingly, platinum compound **2** does not exhibit any reactivity toward phenylacetylene even after heating at 80 °C for 48 h (Scheme 4a).

Scheme 4. Selected Experiments Carried out to Provide Information about the Mechanism of Phenylacetylene Activation Mediated by TMOFLPs 1:2



turn, compounds **1** promote $\text{C}(\text{sp})-\text{H}$ cleavage of the alkyne to form the corresponding digold σ,π -acetylide complexes **11**, but at a considerably slower pace (24 h at 25 °C: **1a**, 9% of **11a**; **1b**, 70% of **11b**; **1c**, 12% of **11c**; NMR spectroscopic yields). To account for the origin of **10**, we performed the reaction of **2** with equivalent amounts of phenylacetylene and catalytic quantities of **1**. In addition, we also tested the same transformation but using **6** as the catalyst, since this presumably forms *in situ* after protonation of **2** by small amounts of HNTf_2 derived from the

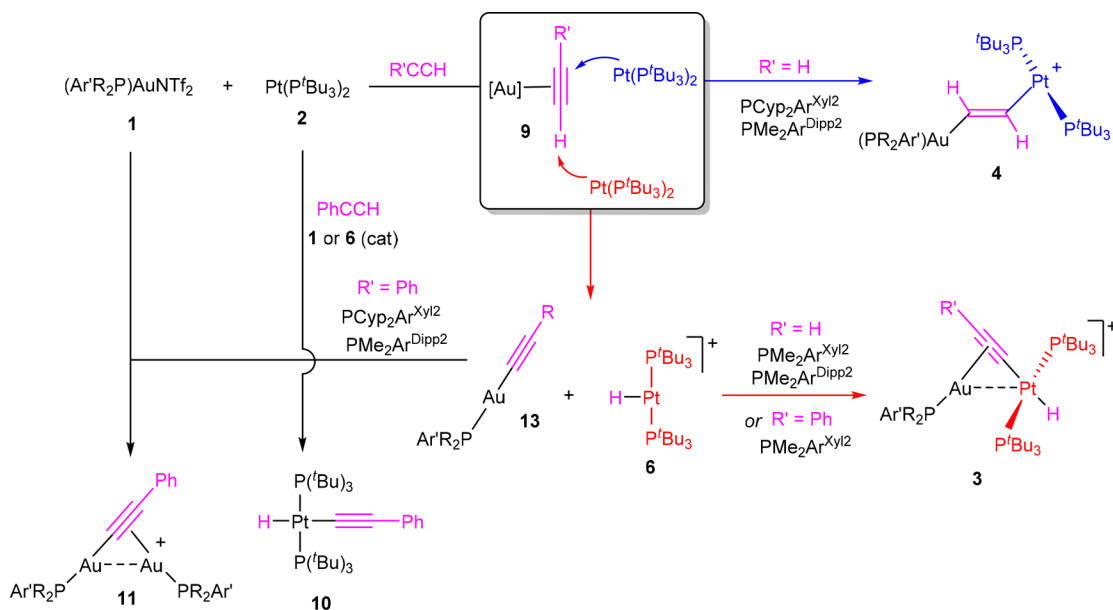
reaction of **1** and phenylacetylene toward **11**. Catalytic amounts of compound **6** may also derive from the cooperative Au/Pt $\text{C}(\text{sp})-\text{H}$ activation of phenylacetylene by a deprotonation mechanism. As anticipated, immediate conversion of **2** into **10** was observed at room temperature using either **1** or **6** in catalytic amounts as low as 2 mol % (Scheme 4b). The reaction of the independently prepared **10** with **1a** leads to **3a** as the major species by NMR analysis, though other unidentified side products are also formed. The reaction of the independently synthesized neutral σ -acetylide compound $[(\text{PMe}_2\text{Ar}^{\text{Dipp}2})\text{Au}(\text{C}\equiv\text{CPh})]$ (**13b**; ORTEP diagram in Figure S4) with **1b** yielded the expected digold σ,π -acetylide **11b** under mild conditions (Scheme 4c). This result parallels the reactivity previously described for NHC-based gold complexes, where the latter transformation proceeds smoothly,^{24f} as also occurs with the parent unsubstituted acetylide ($\text{C}\equiv\text{CH}$) fragment.¹⁸ Thus, terminal acetylides of type **13** may be regarded as key intermediates toward compounds **11** during the activation of PhCCH .

On consideration of all the information discussed above, Scheme 5 contains an overall mechanistic picture to account for the phosphine-dependent product distribution during the activation of alkynes. The common gold acetylene adduct **9** is proposed to be a key intermediate. While both deprotonation and 1,2-addition mechanisms (blue and red in Scheme 5, respectively) are viable for acetylene, only the deprotonation pathway seems to be operative in the case of the more acidic phenyl-substituted alkyne. Once the latter is deprotonated to form an equimolar mixture of **13** and **6**, steric factors appear to dominate the final product distribution. The combination of the more hindered terphenylphosphines with the bulkier phenylacetylene prevents formation of the corresponding Au/Pt heterobimetallic adducts likely due to a steric clash, while the latter is the only observed complex in the $\text{PMe}_2\text{Ar}^{\text{Xyl}2}$ system. In turn, terminal gold acetylides **13**, although unable to react with $[\text{Pt}(\text{P}^t\text{Bu}_3)_2\text{H}]^+$ (**6**), rapidly yield the corresponding digold σ,π -acetylides **11** upon combination with still unreacted triflimide complexes **1**. We can also infer from our experimental observations that several of the transformations depicted in Scheme 5 constitute dynamic equilibria that are dependent on reaction conditions.

CONCLUSIONS

In summary, we provide evidence for the potential of a cooperative Au(I)/Pt(0) bimetallic system that behaves as a metal-only frustrated Lewis pair for the activation of small molecules. We have demonstrated that subtle modifications of the phosphine ligands bound to gold have a strong effect on the regioselectivity of acetylene activation. Thus, simply by adjusting the steric hindrance of those phosphines, we have been able to select the operating mechanism (deprotonation vs 1,2-addition) during acetylene activation with full specificity. This is possible without the need to alter the basicity of the platinum(0) moiety, thus contrasting with main-group FLP systems. Moreover, ligand modification also permitted us to diverge from the product distribution that results from $\text{C}-\text{H}$ bond cleavage in phenylacetylene. While the less congested system based on $\text{PMe}_2\text{Ar}^{\text{Xyl}2}$ yields a heterobimetallic σ,π -acetylide compound, the more hindered pairs constructed around $\text{PMe}_2\text{Ar}^{\text{Dipp}2}$ and $\text{PCyp}_2\text{Ar}^{\text{Xyl}2}$ resulted in the formation of equimolar mixtures of digold σ,π -acetylide compounds, $[\text{Pt}(\text{P}^t\text{Bu}_3)_2(\text{H})]^+$, and $[\text{Pt}(\text{P}^t\text{Bu}_3)_2(\text{H})(\text{C}\equiv\text{CPh})]$. These results highlight one of the advantages of incorporating

Scheme 5. Overall Representation of the Proposed Pathways to Account for Product Distribution during Alkyne Activation



transition metals into frustrated designs, namely the ability to easily tune the stereoelectronic properties of the acidic site, in turn a challenging task in traditional FLPs that rely on the electrophilicity of fluorinated boranes. The straightforward control of these properties in TMFLPs may be exploited for the development of more selective catalytic transformations drawn on the concept of frustration.

EXPERIMENTAL SECTION

General Considerations. All preparations and manipulations were carried out using standard Schlenk and glovebox techniques, under an atmosphere of argon and of high-purity nitrogen, respectively. All solvents were dried, stored over 4 Å molecular sieves, and degassed prior to use. Toluene (C₇H₈) and *n*-pentane (C₅H₁₂) were distilled under nitrogen over sodium. Tetrahydrofuran (THF) and diethyl ether were distilled under nitrogen over sodium/benzophenone. [D₆]-Benzene was dried over molecular sieves (4 Å) and CD₂Cl₂ over CaH₂ and distilled under argon. [AuCl(THT)] (THT = tetrahydrothiophene)²⁵ and compounds **1a**,⁷ **1b**,¹⁸ **1c**,⁷ **2**,²⁶ and **6**¹⁶ were prepared as described previously. Other chemicals were commercially available and were used as received. Solution NMR spectra were recorded on Bruker AMX-300, DRX-400, and DRX-500 spectrometers. Spectra were referenced to external SiMe₄ (δ 0 ppm) using the residual proton solvent peaks as internal standards (¹H NMR experiments) or the characteristic resonances of the solvent nuclei (¹³C NMR experiments), while ³¹P was referenced to H₃PO₄. Spectral assignments were made by routine one- and two-dimensional NMR experiments where appropriate (Figure 5). Infrared spectra were recorded on a Bruker Vector 22 spectrometer, and sampling preparation was carried out in Nujol. For elemental analyses a LECO TruSpec CHN elementary analyzer was utilized. 1965525, 1965526, and 1986501–1986504 contain the supplementary crystallographic data for this paper. These data can be obtained free of charge from the Cambridge Crystallographic Data Centre.

[(PMe₂Ar^{Xyl2})Au(μ-η²-η¹-C≡CH)Pt(H)(P'Bu₃)₂]NTf₂ (3a**).** A solid mixture of **1a** (100 mg, 0.121 mmol) and **2** (73 mg, 0.121 mmol) was dissolved in 5 mL of toluene and stirred at room temperature for 12 h under a C₂H₂ atmosphere (0.5 bar). The solution was layered with pentane (10 mL) and stored at −30 °C overnight to yield compound **3a** as colorless crystals (115 mg, 66%). Anal. Calcd for C₅₂H₈₃AuF₆NO₄P₃PtS₂: C, 43.1; H, 5.8; N, 1.0; S, 4.4. Found: C, 43.1; H, 5.5; N, 1.1; S, 4.4. ¹H NMR (400 MHz, C₆D₆, 25 °C): δ 7.26 (t, 2 H, ³J_{HH} = 7.6 Hz, H_b), 7.18 (m, 1 H, H_d), 7.15 (d, 4 H, ³J_{HH} = 7.6 Hz,

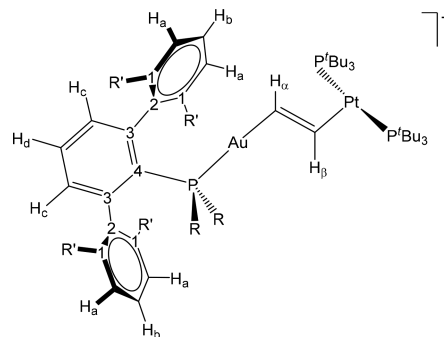


Figure 5. General labeling scheme used for ¹H and ¹³C{¹H} NMR assignments.

H_a), 6.71 (dd, 2 H, ³J_{HH} = 7.5 Hz, ⁴J_{HP} = 3.5 Hz, H_c), 2.98 (m, 1 H, C≡CH), 2.15 (s, 12 H, Me_{Xyl}), 1.52 (d, 6 H, ²J_{HP} = 9.5 Hz, PMe₂), 1.49 (vt, 54 H, ³J_{HP} = 6.4 Hz, ^tBu), −10.27 (m, 1H, ²J_{HP} = 14 Hz, ¹J_{HPt} = 571 Hz, Pt–H). ¹³C{¹H} NMR (100 MHz, C₆D₆, 25 °C): δ 146.8 (d, ²J_{CP} = 10 Hz, C₃), 141.1 (d, ³J_{CP} = 4 Hz, C₂), 136.8 (C₁), 132.4 (CH_d), 131.4 (d, ³J_{CP} = 8 Hz, CH_c), 129.3 (CH_b), 128.8 (CH_a), 127.4 (d, ¹J_{CP} = 57 Hz, C₄), 126.6 (d, ²J_{CP} = 79 Hz, C≡CH), 121.4 (q, ¹J_{CF} = 323 Hz, CF₃), 108.2 (d, ²J_{CP} = 20 Hz, C≡CH), 41.2 (vt, ¹J_{CP} = 8 Hz, Pt–C(CH₃)₃), 33.1 (Pt–C(CH₃)₃), 21.8 (Me_{Xyl}), 17.0 (d, ¹J_{CP} = 38 Hz, PMe₂). ³¹P{¹H} NMR (160 MHz, C₆D₆, 25 °C): δ 82.9 (¹J_{Pt} = 2768 Hz, P'(Bu)₃), 0.8 (Au–P). IR (Nujol): ν(C≡C–H) 3171, ν(C≡C) 1843 cm^{−1}.

[(PCyp₂Ar^{Xyl2})Au(μ-η²-η¹-HC=CH)Pt(P'Bu₃)₂]NTf₂ (4c**).** A mixture of compounds **1c** (100 mg, 0.107 mmol) and **2** (64 mg, 0.107 mmol) was dissolved in toluene (5 mL), and the argon atmosphere was replaced by C₂H₂ (0.5 bar), upon which the bright yellow solution changed to an intense orange. The solution was then filtered, layered with pentane, and stored at −20 °C overnight to yield compound **4c** as orange crystals (68 mg, 41%). Anal. Calcd for C₆₀H₉₅AuF₆NO₄P₃PtS₂: C, 46.3; H, 6.1; N, 1.0; S, 4.1. Found: C, 45.9; H, 5.7; N, 1.0; S, 3.8. ¹H NMR (400 MHz, C₆D₆, 25 °C): δ 7.19–7.08 (m, 3 H, H_b, H_d), 7.01 (d, 4 H, ³J_{HH} = 7.6 Hz, H_a), 6.66 (dd, 2 H, ³J_{HH} = 7.7 Hz, ⁴J_{HP} = 2.7 Hz, H_c), 4.51 (dt, 1 H, ³J_{HH} = 7.6 Hz, ³J_{HP} = 3.4 Hz, ²J_{HPt} = 110 Hz, H_f), 4.20 (dd, 1 H, ³J_{HH} = 7.6 Hz, ³J_{HP} = 6.4 Hz, ³J_{HPt} = 194 Hz, H_g), 2.22–2.09 (m, 2 H, CH), 1.97 (s, 24 H, Me_{Xyl}), 1.71–1.47 (m, 12 H, CH₂), 1.36 (vt, 54 H, ³J_{HP} = 6.5 Hz, ^tBu), 1.24–1.09 (m, 4H, CH₂). ¹³C{¹H} NMR (100 MHz, C₆D₆, 25 °C): δ 159.4 (d, ²J_{CP} = 113, ¹J_{CH} = 198 Hz, CH_a),

148.9 (d, $^2J_{CP} = 10$ Hz, C₃), 142.4 (d, $^3J_{CP} = 5$ Hz, C₂), 137.9 (C₁), 137.0 (d, $^3J_{CP} = 6$ Hz, CH_c), 132.4 (d, $^4J_{CP} = 8$ Hz, CH_d), 131.6 (d, $^1J_{CP} = 40$ Hz, C₄), 129.3 (CH_c), 125.7 (CH_b), 121.7 (q, $^1J_{CF} = 322$ Hz, CF₃), 115.5 (d, $^3J_{CP} = 24$, $^1J_{CH} = 193$ Hz, CH_β), 41.2 (vt, $^1J_{CP} = 8$ Hz, Pt-C(CH₃)₃), 38.0 (d, $^1J_{CP} = 28$ Hz, PCH), 35.0 (d, $^2J_{CP} = 10$ Hz, CH₂), 33.1 (Pt-C(CH₃)₃), 32.5 (d, $^2J_{CP} = 10$ Hz, CH₂), 26.0 (d, $^2J_{CP} = 10$ Hz, CH₂), 25.7 (d, $^2J_{CP} = 10$ Hz, CH₂), 21.4 (Me_{Xyl}). $^{31}\text{P}\{^1\text{H}\}$ NMR (160 MHz, C₆D₆, 25 °C): δ 70.6 ($^1J_{\text{PPr}} = 3323$ Hz, P(^tBu)₃), 51.7 ($^4J_{\text{PPr}} = 277$ Hz, Au–P). IR (Nujol): $\nu(\text{C}=\text{H})$ 3172, $\nu(\text{C}=\text{C})$ 1645 cm⁻¹.

[(PCyp₂Ar^{Xyl})Au(C≡CH)] (5c). A suspension of (PCyp₂Ar^{Xyl})-AuCl⁷ (200 mg, 0.29 mmol) in toluene (10 mL) was cooled to -40 °C, and a toluene solution containing a small excess (1.2 equiv) of Mg(C≡CH)Br was added dropwise. The mixture was stirred for an additional 2 h at -40 °C. The volatiles were removed in vacuo, and the residue was extracted with pentane. Evaporation of the solvent led to compound **5c** as a white powder (35 mg, 18%). Gold acetylides are potentially explosive and should be handled with caution. Anal. Calcd for C₃₄H₄₀AuP: C, 60.4; H, 6.0. Found: C, 60.4; H, 5.8. ^1H NMR (400 MHz, C₆D₆, 25 °C): δ 7.29 (t, 2 H, $^3J_{\text{HH}} = 7.6$ Hz, H_b), 7.09 (d, 4 H, $^3J_{\text{HH}} = 7.56$ Hz, H_a), 6.94 (td, 1 H, $^3J_{\text{HH}} = 7.6$ Hz, $^5J_{\text{HP}} = 1.6$ Hz, H_d), 6.64 (dd, 2 H, $^3J_{\text{HH}} = 7.6$ Hz, $^4J_{\text{HP}} = 2.7$ Hz, H_c), 2.17–2.05 (m, 2 H, PCH), 1.97 (s, 12 H, Me_{Xyl}), 1.78 (d, 1H, $^4J_{\text{HP}} = 5.5$ Hz, AuC≡CH), 1.7–1.1 (m, 16 H, CH₂). $^{13}\text{C}\{^1\text{H}\}$ NMR (100 MHz, C₆D₆, 25 °C): δ 148.8 (d, $^2J_{CP} = 10$ Hz, C₃), 141.9 (d, $^3J_{CP} = 5$ Hz, C₂), 136.3 (C₁), 132.0 (d, $^3J_{CP} = 7$ Hz, CH_c), 130.8 (d, $^4J_{CP} = 5$ Hz, CH_d), 130.2 (d, $^2J_{CP} = 158$ Hz, AuC≡CH), 128.9 (d, $^1J_{CP} = 48$ Hz, C₄), 128.7 (CH_b), 128.4 (CH_a), 86.6 (d, $^3J_{CP} = 25$ Hz, $^1J_{CH} = 233$ Hz, AuC≡CH), 38.3 (d, $^1J_{CP} = 31$ Hz, PCH), 35.1 (d, $^2J_{CP} = 10$ Hz, CH₂), 32.4 (d, $^2J_{CP} = 10$ Hz, CH₂), 25.4 (d, $^2J_{CP} = 11$ Hz, CH₂), 25.3 (d, $^2J_{CP} = 13$ Hz, CH₂), 21.6 (Me_{Xyl}). $^{31}\text{P}\{^1\text{H}\}$ NMR (160 MHz, C₆D₆, 25 °C): δ 55.5. IR (Nujol): $\nu(\text{C}=\text{H})$ 3286, $\nu(\text{C}=\text{C})$ 1944 cm⁻¹.

[Pt(^tBu)₂(H)(C≡CPh)] (10). A toluene (5 mL) solution of **2** (64 mg, 0.107 mmol), phenylacetylene (11 μL , 0.107 mmol), and **1b** or **1c** (2 mg, 0.002 mmol) was stirred at room temperature for 15 min. The volatiles were removed in vacuo, and the residue was extracted with pentane (3 \times 5 mL). Evaporation of the solvent yielded compound **10** as a colorless oil (26 mg, 35%). Suitable crystals for X-ray diffraction studies can be obtained by slow pentane evaporation at room temperature. Anal. Calcd for C₃₂H₆₀Pt: C, 54.8; H, 8.6. Found: C, 54.9; H, 8.4. ^1H NMR (400 MHz, C₆D₆, 25 °C): δ 7.64 (d, 2 H, $^3J_{\text{HH}} = 7.3$ Hz, *o*-C₆H₅), 7.21 (t, 2 H, $^3J_{\text{HH}} = 8.2$ Hz, *m*-C₆H₅), 7.02 (t, 1 H, $^3J_{\text{HH}} = 7.3$ Hz, *p*-C₆H₅), 1.58 (vt, 54 H, $^3J_{\text{HP}} = 6.3$ Hz, ^tBu), -9.46 (t, 1 H, $^2J_{\text{HP}} = 15.4$, $^1J_{\text{HPt}} = 532.9$ Hz, Pt-H). $^{13}\text{C}\{^1\text{H}\}$ NMR (100 MHz, C₆D₆, 25 °C): δ 131.3 (CH-Ar), 130.7 (CH-Ar), 129.3 (C-Ar), 124.2 (CH-Ar), 118.8 (C≡CPh), 117.6 (C≡CPh), 40.4 (vt, $^1J_{CP} = 8$ Hz, $^2J_{\text{Cpt}} = 35$ Hz, Pt-P(C(CH₃)₃), 33.1 (Pt-P(C(CH₃)₃)). $^{31}\text{P}\{^1\text{H}\}$ NMR (160 MHz, C₆D₆, 25 °C): δ 81.3 ($^1J_{\text{PPr}} = 2880$ Hz). IR (Nujol): $\nu(\text{C}=\text{C})$ 2090 cm⁻¹.

[(PR₂Ar)₂Au₂(μ - η^2 -C≡CPh)]NTf₂ (11). To a solution of compounds **1** (100 mg, 0.107 mmol) and **2** (64 mg, 0.107 mmol) in toluene (5 mL) was added 1 equiv of phenylacetylene (11 μL , 0.107 mmol) and the mixture stirred at room temperature for 15 min. The volatiles were removed in vacuo, and the residue was washed with pentane (3 \times 5 mL). The resulting fine white powder (60 mg, **11b**; 65 mg, **11c**) contains a mixture of compounds **11** and **6**, the latter of which could not be separated by common methods but whose spectroscopic features did not hamper full characterization of compounds **11**. A single crystal of **11b** suitable for X-ray diffraction studies was grown by slow diffusion of pentane into a toluene solution (2:1 v/v) at -30 °C. Alternatively, compound **11b** could be synthesized free of **6**: to a solution of **1b** (100 mg, 0.107 mmol) in toluene (5 mL) was added 1 equiv of phenylacetylene (11 μL , 0.107 mmol) and the mixture stirred at room temperature for 18 h. The volatiles were removed in vacuo, and the residue was washed with pentane (3 \times 5 mL) to yield **11b** as a white powder (90 mg, 50%). Data for compound **11b** are as follows. ^1H NMR (400 MHz, C₆D₆, 25 °C): δ 7.55 (d, 2 H, $^3J_{\text{HH}} = 7.6$ Hz, *o*-C₆H₅), 7.35 (m, 2 H, *m*-C₆H₅), 7.26 (t, 4 H, $^3J_{\text{HH}} = 7.6$ Hz, H_b), 7.09 (d, 8 H, $^3J_{\text{HH}} = 7.8$ Hz, H_a), 7.02 (t, 1 H, $^3J_{\text{HH}} = 8.4$ Hz, *p*-C₆H₅), 6.93 (m, 6 H, H_c, H_d), 2.54 (sept, 8 H, $^3J_{\text{HH}} = 6.7$ Hz, ⁱPr(CH)), 1.57 (d, 12 H, $^2J_{\text{HP}} = 9.8$ Hz, PMe₂), 1.29 (d, 24 H, $^3J_{\text{HH}} = 6.7$ Hz, ⁱPr(CH₃)), 0.90 (d, 24 H, $^3J_{\text{HH}} =$

6.7 Hz, ⁱPr(CH₃)). $^{13}\text{C}\{^1\text{H}\}$ NMR (100 MHz, C₆D₆, 25 °C): δ 146.6 (C₁), 146.0 (d, $^2J_{CP} = 10$ Hz, C₃), 138.6 (d, $^3J_{CP} = 3$ Hz, C₂), 134.5 (d, $^3J_{CP} = 9$ Hz, CH_c), 133.0 (CH-Ar), 131.5 (d, $^1J_{CP} = 43$ Hz, C₄), 130.7 (CH-Ar), 130.1 (CH_b), 129.7 (C-Ar), 129.3 (CH_d), 125.7 (CH-Ar), 124.0 (CH_a), 120.9 (q, $^1J_{CF} = 323$ Hz, CF₃), 119.9 (C≡CPh), 116.4 (C≡CPh), 31.6 (ⁱPr(CH)), 25.5 (ⁱPr(CH₃)), 23.1 (ⁱPr(CH₃)), 17.2 (d, $^1J_{CP} = 38$ Hz, PMe₂). $^{31}\text{P}\{^1\text{H}\}$ NMR (160 MHz, C₆D₆, 25 °C): δ 0.4. IR (Nujol): $\nu(\text{C}=\text{C})$ 2048 cm⁻¹. Data for compound **11c** are as follows. ^1H NMR (400 MHz, C₆D₆, 25 °C): δ 7.42 (d, 2 H, $^3J_{\text{HH}} = 8$ Hz, *o*-C₆H₅), 7.24 (t, 4 H, $^3J_{\text{HH}} = 7.6$ Hz, H_b), 7.08 (d, 8 H, $^3J_{\text{HH}} = 7.6$ Hz, H_a), 7.05 (m, 5 H, $^3J_{\text{HH}} = 7.6$ Hz, H_c, *m*-C₆H₅, *p*-C₆H₅), 6.63 (dd, 4 H, $^3J_{\text{HH}} = 7.6$ Hz, $^4J_{\text{HP}} = 3.3$ Hz, H_c), 1.94 (s, 24 H, Me_{Xyl}), 2.25–2.13 (m, 4 H, PCH), 1.87–1.22 (m, 32 H, CH₂). $^{13}\text{C}\{^1\text{H}\}$ NMR (100 MHz, C₆D₆, 25 °C): δ 148.1 (d, $^2J_{CP} = 10$ Hz, C₃), 141.5 (d, $^3J_{CP} = 5$ Hz, C₂), 136.8 (C₁), 133.6 (d, $^3J_{CP} = 7$ Hz, CH_c), 132.5 (d, $^4J_{CP} = 2$ Hz, CH_d), 131.9 (CH-Ar), 130.6 (CH-Ar), 129.3 (CH_b), 129.1 (CH_a), 128.8 (C-Ar), 126.5 (d, $^1J_{CP} = 46$ Hz, C₄), 125.7 (CH-Ar), 124.2 (C≡CPh), 121.6 (C≡CPh), 120.1 (q, $^1J_{CF} = 323$ Hz, CF₃), 40.2 (d, $^1J_{CP} = 36$ Hz, PCH), 38.7 (d, $^3J_{CP} = 7$ Hz, CH₂), 38.4 (d, $^3J_{CP} = 6.6$ Hz, CH₂), 35.8 (d, $^2J_{CP} = 12$ Hz, CH₂), 32.9 (d, $^2J_{CP} = 13$ Hz, CH₂), 21.6 (Me_{Xyl}). $^{31}\text{P}\{^1\text{H}\}$ NMR (160 MHz, C₆D₆, 25 °C): δ 53.9. IR (Nujol): $\nu(\text{C}=\text{C})$ 2029 cm⁻¹.

[(PMe₂Ar^{Xyl})Au(μ - η^2 -C≡CPh)Pt(H)(^tBu)₂]NTf₂ (12a). To a solution of **1a** (100 mg, 0.121 mmol) and **2** (73 mg, 0.121 mmol) in toluene (5 mL) was added 1 equiv of phenylacetylene (13 μL , 0.121 mmol) and the mixture stirred at room temperature for 15 min. The volatiles were removed in vacuo, and the residue was washed with pentane (3 \times 5 mL) to yield compound **12a** as a white powder (70 mg, 38%). Anal. Calcd for C₅₈H₈₇AuF₆NO₄P₃PtS₂: C, 45.7; H, 5.8; N, 0.9; S, 4.2. Found: C, 45.9; H, 5.6; N, 1.1; S, 4.3. ^1H NMR (400 MHz, C₆D₆, 25 °C): δ 7.39 (d, 2 H, $^3J_{\text{HH}} = 7.3$ Hz, *o*-C₆H₅), 7.15 (m, 3 H, *m*-C₆H₅, *p*-C₆H₅), 7.09 (m, 5 H, H_b, H_c, H_d), 6.99 (d, 4 H, $^3J_{\text{HH}} = 7.6$ Hz, H_a), 6.58 (dd, 2 H, $^3J_{\text{HH}} = 7.6$ Hz, $^4J_{\text{HP}} = 3.3$ Hz, H_c), 2.06 (s, 12 H, Me_{Xyl}), 1.38 (vt, 54 H, $^3J_{\text{HP}} = 6.3$ Hz, ^tBu), 1.33 (d, 6 H, $^2J_{\text{HP}} = 10$ Hz, PMe₂), -10.40 (m, 1H, $^2J_{\text{HP}} = 14$ Hz, $^1J_{\text{HPt}} = 608$ Hz, Pt-H). $^{13}\text{C}\{^1\text{H}\}$ NMR (100 MHz, C₆D₆, 25 °C): δ 145.9 (d, $^2J_{CP} = 10$ Hz, C₃), 140.5 (d, $^3J_{CP} = 4$ Hz, C₂), 137.9 (C₁), 136.6 (CH-Ar), 132.2 (d, $^4J_{CP} = 2$ Hz, CH_d), 131.2 (CH-Ar), 129.3 (d, $^3J_{CP} = 9$ Hz, CH_c), 129.2 (C-Ar), 128.9 (CH_b), 128.6 (CH_a), 125.7 (d, $^1J_{CP} = 60$ Hz, C₄), 124.2 (CH-Ar), 121.3 (q, $^1J_{CF} = 323$ Hz, CF₃), 91.1 (C≡CPh), 85.9 (C≡CPh), 41.0 (vt, $^1J_{CP} = 8$ Hz, $^2J_{\text{Cpt}} = 35$ Hz, Pt-P(C(CH₃)₃), 33.1 (Pt-P(C(CH₃)₃), 32.2 (Me_{Xyl}), 21.9 (d, $^1J_{CP} = 36$ Hz, PMe₂). $^{31}\text{P}\{^1\text{H}\}$ NMR (160 MHz, C₆D₆, 25 °C): δ 82.2 ($^1J_{\text{PPr}} = 2810$ Hz, P(^tBu)₃), 3.3 (Au–P). IR (Nujol): $\nu(\text{C}=\text{C})$ 1982 cm⁻¹.

[(PMe₂Ar^{Dipp})Au(C≡CPh)] (13b). Following a previously reported method,²⁷ to a solution of phenylacetylene (36 μL , 0.325 mmol) and KOH (18 mg, 0.325 mmol) in 15 mL of methanol was added a suspension of (PMe₂Ar^{Dipp})AuCl (150 mg, 0.216 mmol). The solution was stirred for 20 h at room temperature, and the solid was filtered and washed with diethyl ether (2 \times 5 mL) to yield compound **13b** as a white solid (139 mg, 85%). Gold acetylides are potentially explosive and should be handled with caution. Anal. Calcd for C₄₀H₄₈AuP: C, 63.5; H, 6.4. Found: C, 63.2; H, 6.7. ^1H NMR (400 MHz, C₆D₆, 25 °C): δ 7.74 (d, 2 H, $^3J_{\text{HH}} = 7.3$ Hz, *o*-C₆H₅), 7.36 (t, 2 H, $^3J_{\text{HH}} = 7.8$ Hz, *m*-C₆H₅), 7.17 (d, 4 H, $^3J_{\text{HH}} = 7.8$ Hz, H_a), 7.06 (t, 1 H, $^3J_{\text{HH}} = 8.3$ Hz, *p*-C₆H₅), 6.99 (m, 5 H, H_b, H_c, H_d), 2.64 (sept, 4 H, $^3J_{\text{HH}} = 6.5$ Hz, ⁱPr(CH)), 1.34 (d, 12 H, $^3J_{\text{HH}} = 6.5$ Hz, ⁱPr(CH₃)), 0.92 (d, 12 H, $^3J_{\text{HH}} = 6.5$ Hz, ⁱPr(CH₃)), 0.90 (d, 6 H, $^2J_{\text{HP}} = 10$ Hz, PMe₂). $^{13}\text{C}\{^1\text{H}\}$ NMR (100 MHz, C₆D₆, 25 °C): δ 146.6 (CH-Ar), 145.7 (d, $^2J_{CP} = 11$ Hz, C₃), 139.0 (d, $^3J_{CP} = 4$ Hz, C₂), 135.8 (C₁), 134.3 (CH-Ar), 132.8 (d, $^3J_{CP} = 7$ Hz, CH_c), 132.6 (d, $^4J_{CP} = 2$ Hz, CH_d), 130.4 (d, $^1J_{CP} = 47$ Hz, C₄), 129.8 (CH_b), 128.9 (CH_a), 125.9 (C-Ar), 123.8 (CH-Ar), 123.0 (C≡CPh), 102.4 (C≡CPh), 31.6 (ⁱPr(CH)), 25.7 (ⁱPr(CH₃)), 23.2 (ⁱPr(CH₃)), 16.9 (d, $^1J_{CP} = 35$ Hz, PMe₂). $^{31}\text{P}\{^1\text{H}\}$ NMR (160 MHz, C₆D₆, 25 °C): δ 9.7. IR (Nujol): $\nu(\text{C}=\text{C})$ 2115 cm⁻¹. MS-ESI (*m/z*): calcd for C₄₀H₄₈AuPt, 756.76; found, [M + H] 757.3, [M + Na] 779.3, [M + K] 795.3.

Computational Details. Geometry optimization of minima and transition states was carried out with the Gaussian software package.²⁸ Optimizations were carried out without symmetry restrictions using the

ω B97xD functional,²⁹ which includes empirical dispersion corrections.³⁰ The 6-31g(d,p) basis set³¹ was used for nonmetal atoms, and Au and Pt atoms were described with the SDD basis and associated electron core potential (ECP).³² Bulk solvent effects (dichloromethane) were included during optimization with the SMD continuum model.³³ Vibrational analysis was carried out on the stationary points to characterize them as minima or transition states as well as to calculate the zero-point corrections, and thermal corrections were made to the enthalpy and free energy. Free energies were corrected (ΔG_{qh}) to account for errors associated with the harmonic oscillator approximation. Thus, according to Truhlar's quasi harmonic approximation, all vibrational frequencies below 100 cm^{-1} were set to this value so that the entropy contribution was not overestimated.³⁴ These anharmonic corrections were calculated with the Goodvibes code.³⁵ The CYLview visualization software has been used to create some of the figures.³⁶

■ ASSOCIATED CONTENT

SI Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/acs.organomet.0c00330>.

Additional computational details, X-ray structural data, and NMR spectra of new compounds (PDF)

Cartesian coordinates of the calculated structures (XYZ)

Accession Codes

CCDC 1965525, 1965526 and 1986501–1986504 contain the supplementary crystallographic data for this paper. These data can be obtained free of charge via www.ccdc.cam.ac.uk/data_request/cif, or by emailing data_request@ccdc.cam.ac.uk, or by contacting The Cambridge Crystallographic Data Centre, 12 Union Road, Cambridge CB2 1EZ, UK; fax: +44 1223 336033.

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Notes

The authors declare no competing financial interest.

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